

Chemistry Worksheets Class 11 on Chapter 4 Chemical Bonding and Molecular Structure - Set 4

Q-1: The compound's formula is X_2Y_3 . How many electrons are in X's and Y's respective outermost orbits?

- a) 6 and 3
- b) 5 and 6
- c) 5 and 2
- d) 2 and 3

Q-2: In which of the following bond angle is maximum?

- a) NH_3
- b) $SiCl_4$
- c) Br_2
- d) SCl_2

Q-3: The electron adds in one of the following orbitals during the transition from O_2 to O_2^- ion. What is it?

- a) π orbitals
- b) π^* orbitals
- c) σ orbitals
- d) σ^* orbitals

Q-4: Which of the following pairs are isostructural?

- a) XeF_2 , IF_2^-
- b) NH_3 , BF_3
- c) CO_3^{2-} , SO_3^{2-}
- d) PCl_5 , IF_5

Q-5: Which of the following compound is electron deficient?

- a) NCl_3
- b) CH_4
- c) BCl_3
- d) PCl_3

Q-6: In each of the following pairs of compounds, which one is more covalent and why?

- i) $AgCl$, AgI
- ii) CuO , CuS
- iii) $SnCl_2$, $SnCl_4$

Modern Chemistry Chapter 1review Answers Molecular Composition Of Gases

D Kirk



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