Name: ______ Date: _____ Pd:____

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.

$$_1$$
Al₂(SO₄)₃ + $_3$ Ca(OH)₂ \rightarrow $_2$ Al(OH)₃ + $_3$ CaSO₄
342.17 g/mol 74.10 g/mol 78.01 g/mol 136.15 g/mol

STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: W5.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

96 Al =
$$\frac{total\ mass\ Al}{molar\ mass\ Al_2(SO_4)_3}$$
 x 100 % = $\frac{53.96\ g}{342.17\ g}$ x 100 % = 15.77 %

STEP 4: Complete the following problems using a factor chart with units & chemicals. These problems could require ONE ratio. TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

How many grams of calcium sulfate are produced from reacting 2.59 moles of aluminum sulfate?

2.59 mol Al₂(SO₄)₃ x
$$\frac{3 mol CaSO_4}{1 mol Al_2(SO_4)_3}$$
 x $\frac{136 15 g CaSO_4}{1 mol CaSO_4}$ = 1057.89 g CaSO₄

moles A problem B grams B

Modern Chemistry Chapter Mixed Review Stoichiometry Answers

Lingjun Ying

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- 1. Understanding the eBook Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - The Rise of Digital Reading Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Personalized Recommendations
 - Modern Chemistry Chapter Mixed Review Stoichiometry Answers User Reviews and Ratings

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- 5. Accessing Modern Chemistry Chapter Mixed Review Stoichiometry Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Interactive Elements Modern Chemistry Chapter Mixed Review Stoichiometry Answers
- 8. Staying Engaged with Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter Mixed Review Stoichiometry Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter Mixed Review Stoichiometry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Setting Reading Goals Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Fact-Checking eBook Content of Modern Chemistry Chapter Mixed Review Stoichiometry Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web an umbrella term not defined under international law reflecting the common lay understanding of a person who moves away from his or her place of usual residence whether within a country or across an international border temporarily or permanently and for a variety of reasons