Multiple Choice Reaction Mechanisms

23. Given the following reaction mechanism, what is the equation for the overall reaction?

$$2A \rightarrow B + 2C \text{ (slow)}$$

$$B + C \rightarrow D + E (fast)$$

$$C + D \rightarrow E + F(fast)$$

A)
$$2A \rightarrow 2E + F$$
 B) $2A + B + 2C \rightarrow D$
D) $2A + B + 2C + D \rightarrow B + 2C + D + 2E + F$

$$2C \rightarrow D + 2E + F$$

B)
$$2A + B + 2C \rightarrow D + 2E + F$$
 C) $2A + 2C \rightarrow 2E + F$

E)
$$2A + C \rightarrow 2E + F$$

AINS: A

11. Which step below is the rate-determining step?

fasst:

very fast. show

wery fast.

schower.

AND D

ANS: C

12. Which of the following rate law expressions would support the mechanism below?

$$WXZ + Z \rightarrow WXZ, \qquad fast$$

A)
$$r = k[W_1X][Z]$$
 B) $r = k[W]^2$ C) $r = k[WXZ][Z]$ D) $r = k[W]^2[XY]$

ANS: A

11. Consider the following reaction mechanism:

The overall reaction is: A +2B + 300kJ -- C + F

How might the rate of the forward reaction be increased?

ANS: A

Consider the following reaction mechanism:

Step 1:
$$2NO + H_2 \rightarrow N_2 + H_3O_3$$

In this reaction, Ho is ac-

Modern Chemistry Reaction Energy Multiple Choice Answers

Judd E. Hollander

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