

## Molarity and Dilution:

### Molarity:

- When we talk about solutions, oftentimes instead of referring to a number of moles of a solute, we describe concentrations → this is how many moles of the solute are present per 1 litre of solution.
- The unit we use for this is molarity, which means moles of solute per litre of solution.
  - $M = \frac{\text{moles}}{\text{litre}}$
- For example, a 2 molar solution has 2 moles of solute for every litre of solution.
  - $2M = \frac{2 \text{ moles}}{\text{litre}}$
- If we want to know the molarity of a solution and we have the grams of solute, we just convert grams to moles using the molar mass of the substance, and divide by the total volume of the solution.
- Example: 0.100L  $H_2O$  solution with 10.0g NaCl:
  - $10.0\text{g NaCl} \times \frac{1 \text{ mole NaCl}}{58.44\text{g NaCl}} = 0.171 \text{ moles of NaCl}$
  - $\frac{0.171 \text{ mol NaCl}}{0.100\text{L } H_2O} = 1.71 \text{ M NaCl solution}$

### Dilution:

- We can alter the concentration of a solution by a process called dilution. This is the act of adding solvent to a solution. So if the moles of solute stay the same, the solution will become less concentrated as the volume increases.
- We make dilution calculations using this equation:  $M_1V_1 = M_2V_2$
- Plug in the initial concentrations and volume of the solution, then the final volume you will dilute to, then solve for the new concentration.
  - $(1.0M)(1.0L) = M_2(2.0L)$
  - $\text{Molarity} = \frac{\text{moles of solute}}{\text{litres of solution}}$
  - $M_2 = \frac{1 \text{ mol}}{2 \text{ litres}} = 0.50 \text{ molarity (M)}$
- The quantities are inversely related, so doubling the volume will halve the concentration.

### Practice problems:

What would be the molarity of a solution made from 86.5g  $Ca(OH)_2$  in 200mL of water?

- $\text{Molarity} = \frac{\text{moles of solute}}{\text{litres of solution}}$
- $86.5\text{g } Ca(OH)_2 \times \frac{1 \text{ mol } Ca(OH)_2}{74.1\text{g } Ca(OH)_2} = 1.17 \text{ moles}$
- $M = \frac{1.17 \text{ moles}}{0.2\text{L}} = 5.85 \text{ M}$

To what volume would you need to dilute the above solution to make it 1.95 M?

- $M_1V_1 = M_2V_2$
- $(5.85M)(0.2L) = (1.95M) V_2$
- $\frac{5.85}{1.95} = 3$
- $0.2L \times 3 = 0.6L$
- $V_2 = 0.600L$

# Molarity And Dilution Review Document

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