### Molarity and Dilution:

#### Molarity:

- When we talk about solutions, oftentimes instead of referring to a number of moles of a solute, we
  describe concentrations → this is how many moles of the solute are present per 1 litre of solution.
- The unit we use for this is molarity, which means moles of solute per litre of solution.
  - $\odot$   $\mathbf{M} = \frac{moles}{4sre}$
- For example, a 2 molar solution has 2 moles of solute for every litre of solution.
  - ⇒ 2M = 2 moles
    torre
- If we want to know the molarity of a solution and we have the grams of solute, we just convert grams to
  moles using the molar mass of the substance, and divide by the total volume of the solution.
- Example: 0.100L H<sub>2</sub>O solution with 10.0g NaCl:
  - 10.0g NaCl × 1 mole NaCl / 58.44g NaCl = 0.171 moles of NaCl
  - o 0.171 mol NaCl = 1.71 M NaCl solution

#### Dilution:

- We can alter the concentration of a solution by a process called dilution. This is the act of adding solvent
  to a solution. So if the moles of solute stay the same, the solution will become less concentrated as the
  volume increases.
- We make dilution calculations using this equation: M, V, = M, V,
- Plug in the initial concentrations and volume of the solution, then the final volume you will dilute to, then solve for the new concentration.
  - $\circ$  (1.0M)(1.0L) = M<sub>2</sub>(2.0L)
  - Molarity = moles of solute litres of solution
  - $0 \quad \mathbf{M}_2 = \frac{1 \text{ mod}}{2 \text{ targer}} = 0.50 \text{ molarity (M)}$
- The quantities are inversely related, so doubling the volume will halve the concentration.

#### Practice problems:

What would be the molarity of a solution made from 86.5g Ca(OH)2 in 200mL of water?

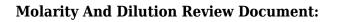
- Molarity = moles of solution
- 86.5g Ca(OH)<sub>2</sub> × 1 mol Ca(OH)<sub>2</sub> / 74.1g Ca(OH)<sub>2</sub> = 1.17 moles
- M = 1.17 moles = 5.85 M

To what volume would you need to dilute the above solution to make it 1.95 M?

- M<sub>1</sub>V<sub>1</sub> = M<sub>2</sub>V<sub>2</sub>
- (5.85M)(0.2L) = (1.95M) V<sub>2</sub>
- $\frac{5.85}{1.95} = 3$
- 0.2L x 3 = 0.6L
- V<sub>2</sub> = 0.600L

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**Hussin A.Rothana** 



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## **Table of Contents Molarity And Dilution Review Document**

- 1. Understanding the eBook Molarity And Dilution Review Document
  - The Rise of Digital Reading Molarity And Dilution Review Document
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Molarity And Dilution Review Document
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molarity And Dilution Review Document
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Molarity And Dilution Review Document
  - Personalized Recommendations
  - Molarity And Dilution Review Document User Reviews and Ratings
  - Molarity And Dilution Review Document and Bestseller Lists
- 5. Accessing Molarity And Dilution Review Document Free and Paid eBooks
  - Molarity And Dilution Review Document Public Domain eBooks
  - Molarity And Dilution Review Document eBook Subscription Services
  - Molarity And Dilution Review Document Budget-Friendly Options
- 6. Navigating Molarity And Dilution Review Document eBook Formats
  - o ePub, PDF, MOBI, and More
  - Molarity And Dilution Review Document Compatibility with Devices
  - Molarity And Dilution Review Document Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molarity And Dilution Review Document
  - Highlighting and Note-Taking Molarity And Dilution Review Document
  - Interactive Elements Molarity And Dilution Review Document

- 8. Staying Engaged with Molarity And Dilution Review Document
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molarity And Dilution Review Document
- 9. Balancing eBooks and Physical Books Molarity And Dilution Review Document
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molarity And Dilution Review Document
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity And Dilution Review Document
  - Setting Reading Goals Molarity And Dilution Review Document
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity And Dilution Review Document
  - Fact-Checking eBook Content of Molarity And Dilution Review Document
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - $\circ \ \ Integration \ of \ Multimedia \ Elements$
  - Interactive and Gamified eBooks

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