

## 2.3 – Mole Conversion – Mole-Mass

Practice # 1 – How many moles are there in 450 grams of  $\text{Na}_2\text{SO}_4$ ? ①

$$450 \text{ g } \text{Na}_2\text{SO}_4 \times \frac{1 \text{ mol } \text{Na}_2\text{SO}_4}{142.04 \text{ g } \text{Na}_2\text{SO}_4} = 3.168091 = 3.2 \times 10^0 \text{ mol } \text{Na}_2\text{SO}_4$$

*Handwritten notes: A red arrow points from 'g' to 'mol'. The final answer is boxed. The unit 'Na<sub>2</sub>SO<sub>4</sub>' is written below the box.*

Practice # 2 – How many grams are there in  $7.4 \times 10^2$  moles of silver?

$$7.4 \times 10^2 \text{ mol } \text{Ag} \times \frac{107.87 \text{ g } \text{Ag}}{1 \text{ mol } \text{Ag}} = 798.238 \times 10^2 = 8.0 \times 10^4 \text{ g } \text{Ag}$$

*Handwritten notes: A red arrow points from 'mol' to 'g'. The final answer is boxed. The unit 'Ag' is written below the box.*

# Mole Ratio Conversion Test Answers

**Tracy Lynn Haroff**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

## **Mole Ratio Conversion Test Answers:**

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