

Mole Ratios

How can the coefficients in a chemical equation be interpreted?

Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers (10^{23} or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

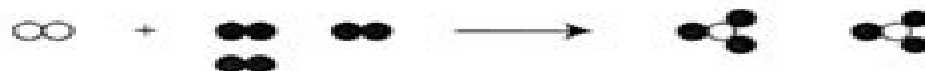
Model 1 – A Chemical Reaction



1. Consider the reaction in Model 1.
 - a. What are the coefficients for each of the following substances in the reaction?



- b. Draw particle models below to illustrate the reaction in Model 1.



2. Consider each situation below as it relates to the reaction in Model 1.
 - a. Calculate the amount of reactants consumed and products made.
 - b. Record the ratio of N_2 to H_2 to NH_3 . Reduce the ratio to the lowest whole numbers possible.

| | N_2 Consumed | H_2 Consumed | NH_3 Produced | Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$ (reduced) |
|---|--------------------------|--------------------------|---------------------------|--|
| For a single reaction, how many molecules of each substance would be consumed or produced? | 1 molecule | 3 molecules | 2 molecules | 1:3:2 |
| If the reaction occurred one hundred times, how many molecules would be consumed or produced? | 100 molecules | 300 molecules | 200 molecules | 1:3:2 |
| If the reaction occurred 538 times, how many molecules would be consumed or produced? | 538 molecules | 1614 molecules | 1076 molecules | 1:3:2 |

Mole Ratio Packet Answer Key

NARAYAN CHANGDER

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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A Mole of Chemistry Caroline Desgranges, Jerome Delhommelle, 2020-03-03 A Mole of Chemistry An Historical and Conceptual Approach to Fundamental Ideas in Chemistry is intended for students in their undergraduate years who need to learn the basics of chemistry including science and engineering as well as humanities This is a companion textbook which provides a unique perspective on how the main scientific concepts describing nature were discovered and eventually how modern chemistry was born The book makes use of context found in history philosophy and the arts to better understand their developments and with as few mathematical equations as possible The focus is then set on scientific reasoning making this book a great companion and addition to traditional chemistry textbooks Features A companion for a general chemistry textbook and provides an historical approach to fundamental chemistry Presents origins of fundamental ideas in chemical science and the focus is then set on scientific reasoning User friendly and with as few mathematical equations as possible About the Authors Dr Caroline Desgranges earned a DEA in Physics in 2005 at the University Paul Sabatier Toulouse III France and a PhD in Chemical Engineering at the University of South Carolina USA in 2008 Dr Jerome Delhommelle earned his PhD in Chemistry at the University of Paris XI Orsay France in 2000 He is currently working as an Associate Professor in Chemistry at the University of North Dakota

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You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards

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