Mole Ratios

How can the coefficients in a chemical equation be interpreted?

Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers (10^{23} or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

Model 1 — A Chemical Reaction

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

1. Consider the reaction in Model 1.

   a. What are the coefficients for each of the following substances in the reaction?

   1 N_2 *3* H_2 *2* NH_3

   b. Draw particle models below to illustrate the reaction in Model 1.

2. Consider each situation below as it relates to the reaction in Model 1.

   a. Calculate the amount of reactants consumed and products made.

   b. Record the ratio of N_2 to H_2 to NH_3. Reduce the ratio to the lowest whole numbers possible.

	N_2 Consumed	H_2 Consumed	NH_3 Produced	Ratio N_2:H_2:NH_3 (reduced)
For a single reaction, how many molecules of each substance would be consumed or produced?	*1 molecule*	*3 molecules*	*2 molecules*	*1:3:2*
If the reaction occurred one hundred times, how many molecules would be consumed or produced?	*100 molecules*	*300 molecules*	*200 molecules*	*1:3:2*
If the reaction occurred 538 times, how many molecules would be consumed or produced?	*538 molecules*	*1614 molecules*	*1076 molecules*	*1:3:2*

Mole Ratios Pogil Answers

NARAYAN CHANGDER

Mole Ratios Pogil Answers:

  Mole Ratios Lifeliqe,2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions   STOICHIOMETRY NARAYAN CHANGDER,2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel https www youtube com smartquiziz I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today s academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more   *You Can Do Chemistry* Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018   **Chemistry Equations & Answers** Speedy Publishing,2014-08-01 One of the most important parts of learning chemistry is simply knowing all of the equations and formulas that are used in it It is important to know the exact formatting of these equations and most classes will require a student to know them for exams It is a good idea to learn these equations and formulas with the use of a study pamphlet The pamphlet can condense all of the information so a student can memorize the equations and formulas while studying

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