

1
→2
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Mole Ratio

$$16.5 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} \times \frac{1 \text{ mol Ba(NO}_3)_2}{2 \text{ mol HNO}_3} \times \frac{261.3 \text{ g}}{1 \text{ mol Ba(NO}_3)_2} = 34.2 \text{ g}$$

use the molar mass of HNO_3 to
calculate its moles

use mole ratio of
 HNO_3 and
 $\text{Ba(NO}_3)_2$ to
calculate the moles
of $\text{Ba(NO}_3)_2$

use moles of
 $\text{Ba(NO}_3)_2$ to
calculate its mass

Mole To Mole Stoichiometric Calculations Answers

Charles H. Corwin



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