

Empirical/Molecular Formula

Name: _____

Date: _____

1. What is an empirical formula? How is it different from a molecular formula?

The empirical formula is the lowest whole number ratio of atoms in a compound. The molecular formula shows the true number of atoms in the molecule and is a whole number multiple of the empirical formula.

2. State the empirical formula for each of the following.



3. A sample of a compound is 50.0% sulfur and 50.0% oxygen by mass. What is the empirical formula?

$$S: \frac{50.0 \text{ g S}}{32.07 \text{ g S}} \times \frac{1 \text{ mol S}}{1} = \frac{1.559 \text{ mol S}}{1.559 \text{ mol}} = 1 \text{ S}$$

EF: SO_2

$$O: \frac{50.0 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{3.125 \text{ mol O}}{1.559 \text{ mol}} = 2 \text{ O}$$

4. A sample of a compound is 52.9% aluminum. The rest is oxygen. What is the empirical formula?

$$Al: \frac{52.9 \text{ g Al}}{26.98 \text{ g Al}} \times \frac{1 \text{ mol Al}}{1} = \frac{1.961 \text{ mol Al}}{1.961 \text{ mol Al}} = 1 \text{ Al} \times 2 = 2$$

EF: Al_2O_3

$$O: \frac{47.1 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{2.944 \text{ mol O}}{1.961 \text{ mol}} = 1.5 \times 2 = 3 \text{ O}$$

Ratio must be a whole

5. A compound is 48.65% carbon, 8.11 % hydrogen and 43.24 % oxygen. Determine the empirical formula.

4.05 mol carbon, 8.11 mol H, 2.70 mol O



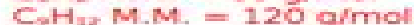
1. Divide by the smallest # to get = $C_{1.5}H_3O_1$

2. Multiply to get a whole number ratio (x2) = $C_3H_6O_2$

6. Vinegar is a dilute solution of acetic acid. The molar mass of acetic acid is 60.0 g/mol and it has an empirical formula of CH_2O . What is the molecular formula of acetic acid?



7. A compound has an empirical formula of C_3H_4 . Which of the following are possible molar masses of the compound? 20 g/mol, 55 g/mol, 80 g/mol, 120 g/mol.



Molecular Formula Worksheet Chemfiesta

John Buckingham



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Formula Index Sadtler Research Laboratories, 1967 Through the Molecular Maze Allen E. Breed, Thomas Rodella, Ronald Basmajian, 1975 Introduces basic chemical principles necessary to the study of biology Assignments in Chemistry Peter Gribben, Wendy St. Germain, 2005 Provides photocopiable worksheets designed to challenge and stimulate students as they explore a variety of chemistry concepts *Molecular formula index* Philadelphia Sadtler Research Laboratories, 1967

Molecular Formula Index, 1987 **Rapid Calculation of Molecular Formulas from Mass Values** Stanford University. Computer Science Dept. Heuristic Programming Project, J. Lederberg, 1972 **Molecular Formula Index**, 1995

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understanding of the discipline through a practical and methodical approach providing solid skills that are essential for continuing in scientific studies **Molecular formula index** American Chemical Society. Chemical Abstracts Service, 1968

Molecular Formula and Structure Angelina R. Sekirnik, 2015 Presented lectures and prepared assessments for graduate training programme in Introduction to Organic chemistry non specialist scientists through the Doctoral Training Centre at the University of Oxford *EINECS: Molecular formula index. pt. 1-2* , 1987

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