

1. What are the four most commonly occurring elements in organisms?

Carbon, Hydrogen, Nitrogen, Oxygen

2. Define organic molecule.

A molecule that contains carbon (with the exception of carbonates - CO_2 and - CO_3)

3. Why is carbon the ideal building block for large molecules?

Carbon can make covalent bond with up to four other atoms

4. Define the following terms:

a. Metabolism:

The overall network of chemical reactions that occur in cells

b. Catabolism:

The set of reactions that break larger molecules down to smaller ones

c. Anabolism:

The set of reactions that assemble large molecules from smaller parts

5. Outline the theory of Vitalism and how Wohler's synthesis of Urea disproved the theory:

Vitalism was a theory that organic compounds could only be made by living things due to their "vital force." When Wohler was able to produce Urea (an organic compound) artificially in his lab it showed that organic compounds could still be formed without the "vital force" of life.

6. Complete the table of common organic molecule groups / parts:

Name of Group	Hydroxyl	Amine	Carboxyl	Methyl
Full Structure	$\text{S}-\text{O}-\text{H}$	$\text{S}-\text{N}-\text{H}$ $\quad \quad \quad \text{H}$	$\text{S}-\text{C}-\text{OH}$ $\quad \quad \quad \text{O}$	$\text{S}-\text{C}-\text{H}$ $\quad \quad \quad \text{H}$ $\quad \quad \quad \text{H}$
Simplified Notation	$-\text{OH}$	$-\text{NH}_2$	$-\text{COOH}$	$-\text{CH}_3$

7. Complete the table of common organic molecules:

Name	Ribose	Glucose	Saturated Fatty Acid	Amino Acid
Drawing				

8. Describe the structure a single water molecule and how this leads to its dipolar (polar) nature:

Two Hydrogen atoms are bonded to one Oxygen atom by covalent bonds. However, electrons in these bonds are not shared evenly so the Oxygen takes on a partial negative charge and Hydrogen takes

9. In the space below, draw three water molecules attracted to one another by hydrogen bonding. Include labels to show the polarity of the molecules and the hydrogen bonds:



a
partial
positive
charge

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