

Specific Heat of a Metal Lab

Objectives:

- Determine the specific heat of a metal.
- Calculate percent error of experimental results

Introduction

Chemists identify substances on the basis of their chemical and physical properties. One physical property of a substance is the amount of energy it will absorb per unit of mass. This property can be measured quite accurately and is called *specific heat* (C_p).

Specific heat is the amount of energy measured in joules, needed to raise the temperature of one gram of the substance one Celsius degree. Often applied to metallic elements, specific heat can be used as a basis for comparing energy absorption and transfer.

To measure specific heat in the laboratory a *calorimeter* of some kind must be used. A calorimeter is a well-insulated container used in measuring energy changes. The calorimeter is insulated to reduce the loss or gain of energy to or from the surroundings. Energy always flows from an object at a higher temperature to an object at a lower temperature. The heat gained by the cooler substance equals the heat lost by the warmer substance, if we assume no loss of heat to the surrounding environment.

$$\text{heat lost} = \text{heat gained}$$

In this experiment, you will determine the specific heat of a metal sample. The metal sample will be heated to a high temperature then placed into a calorimeter containing a known quantity of water at a lower temperature. Having measured the mass of the water in the calorimeter, the temperature change of the water (ΔT) and knowing the specific heat of water ($4.184 \text{ J/g } ^\circ\text{C}$) the heat gained by the water (lost by the metal) can be calculated as follows:

$$Q = (m)(\Delta T)(C)$$

<i>heat gained by the water</i>	=	<i>mass of water (g)</i>	x	<i>change in temperature (ΔT)</i>	x	<i>specific heat of water ($4.184 \text{ J/g } ^\circ\text{C}$)</i>
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The specific heat of the metal can now be calculated:

<i>Specific heat, C =</i>	$\frac{\text{heat gained by the water, } Q}{\text{of metal mass of metal (g) } m \times \Delta T \text{ of metal } (^\circ\text{C})}$
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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Lab 22 Specific Heat Of Metals Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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