

Additivity of Heats of Reaction: Hess's Law

Advanced Study Questions:

Name _____

One of the reactions below is the same as the combination (addition) of the other two reactions. Therefore, according to Hess's Law, the heat of reaction, ΔH , of that reaction should be equal to the sum of the heats of reaction for the other two.

- (1) $\text{NaOH(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$ ΔH_1
- (2) $\text{NaOH(s)} + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O(l)} + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ ΔH_2
- (3) $\text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O(l)} + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ ΔH_3

Determine which of the three reactions above may result from the direct addition of the other two. Clearly show how two equations add up to give the third one. (*Use equations as they are: do not reverse, adjust coefficients, etc.*)

For the reaction obtained as a sum of two others (called the overall reaction), show how can its ΔH be expressed using ΔH for two other reactions? *You do not need values, use symbols only.*

In this experiment, sodium hydroxide is delivered in two ways, as solid or in an aqueous solution. Which sample contains **larger number of moles** of NaOH: 2.65 g solid NaOH or 55.2 mL of 1.15M NaOH solution? *Show full work including units.*

Lab 35 Heats Of Reaction Answers

Michael Seilmaier

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