AP Chemistry Chapter 14 Answers - Zumdahl

14.27

a. $HClO_4(aq) + H_2O(1) \rightarrow H_3O^*(aq) + ClO_4(aq)$. Only the forward reaction is indicated since $HClO_4$ is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $HClO_4(aq) \rightarrow H^*(aq) + ClO_4(aq)$. This reaction is also called the K_a reaction as the equilibrium constant for the reaction is called K_a .

b. Propanioc acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is: $CH_3CH_2CO_2H(aq) + H_2O(l) \Rightarrow H_3O^* + CH_3CH_2CO_2(aq)$ or $CH_3CH_2CO_2H(aq) \Rightarrow$ $H^*(aq) + CH_3CH_2CO^*(aq)$.

NH₄* is a weak acid. Similar to propanoic acid, the dissociation reaction is:

$$NH_4^*(aq) + H_2O(1) \Rightarrow H_3O^*(aq) + NH_3(aq) \text{ or } NH_4^*(aq) \Rightarrow H^*(aq) + NH_3(aq)$$

14.29

An acid is a proton (H*) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H*).

| | Acid | Base | Conjugate Base of Acid | Conjugate Acid of Base |
|----------|--|--|--|--|
| a. b. | H ₂ CO ₃ C ₃ H ₃ NH* C ₄ H ₄ NH* | H ₂ O H ₂ O HCO ₁ | HCO ₃ * C ₃ H ₃ N C ₃ H ₃ N | H ₂ O* H ₂ O* H ₂ CO ₃ |

14.31

Strong acids have a K_a >> 1 and weak acid have K_a < 1. Table 14.2 in the text lists some K_a values for weak acids. K_a values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl, HBr, HI, HNO₃, HClO₄, and H₂SO₄.

- HClO₄ is a strong acid.
- HOCl is a weak acid (K_a = 3.5 x 10⁻⁸).
- e. H₂SO₄ is a strong acid.
- d. H₂SO₃ is a weak diprotic acid with K_{a1} and K_{a2} values less than one.

14.33

The K_a value is directly related to acid strength. As K_a increases, acid strength increases. For water, use K_w when comparing the acid strength of water to other species. The K_a values are:

HClO₄: strong acid (
$$K_a >> 1$$
); HClO₂: $K_a = 1.2 \times 10^{-2}$
NH₄*: $K_a = 5.6 \times 10^{-10}$; H₂O: $K_a = K_w = 1.0 \times 10^{-14}$

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