



7. What is the maximum mass of aluminum hydroxide that can be produced when 25.0 mL of 0.248 M sodium hydroxide solution are mixed with 25.0 mL of 0.189 M aluminum nitrate solution?

Find mol of each reactant

$$0.0250 \text{ L} \times 0.248 \frac{\text{mol}}{\text{L}} = 6.2 \times 10^{-3} \text{ mol NaOH}$$

$$0.0250 \text{ L} \times 0.189 \frac{\text{mol}}{\text{L}} = 4.73 \times 10^{-3} \text{ mol Al}(\text{NO}_3)_3$$

$$6.20 \times 10^{-3} \text{ mol NaOH} \times \frac{1 \text{ mol Al}(\text{OH})_3}{3 \text{ mol NaOH}} \times \frac{78.0 \text{ g Al}(\text{OH})_3}{1 \text{ mol}} =$$

$$\boxed{0.161 \text{ g Al}(\text{OH})_3}$$

need 3 times as much NaOH as  $\text{Al}(\text{NO}_3)_3$  but we only have 1.3 X as much NaOH is limiting

8. Indicate which of the following are redox reactions. For those that are redox reactions, label the reactant that is the oxidizing agent and the reactant that is the reducing agent. You do not have to balance these reactions!



Redox agent agent

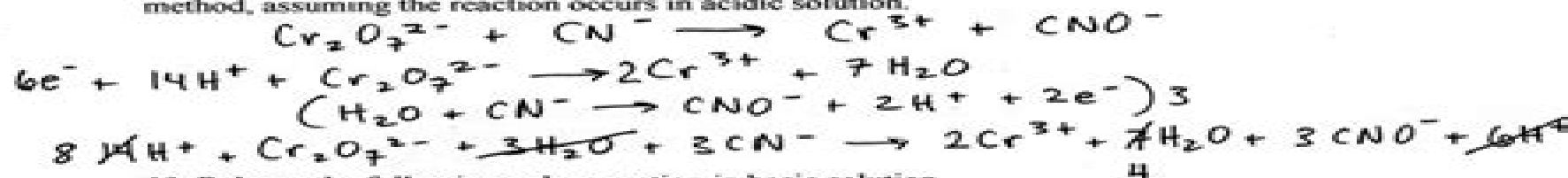


Not redox

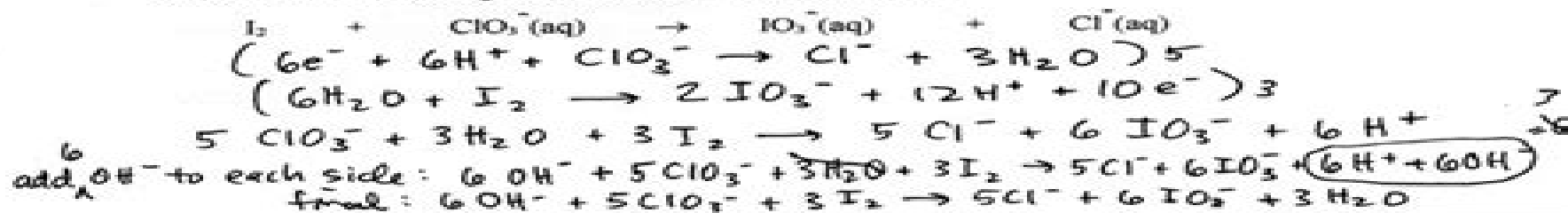


red agent ox agent

9. Aqueous dichromate ion reacts with aqueous cyanide ion ( $\text{CN}^-$ ) to produce chromium (III) ion and aqueous cyanate ( $\text{CNO}^-$ ) ion. Balance the equation using the half-reaction method, assuming the reaction occurs in acidic solution.



10. Balance the following redox reaction in basic solution.



# Modern Chemistry Homework Chapter9 Answers

**Richard Bailey**



## Modern Chemistry Homework Chapter9 Answers:

Essentials of Modern Materials Science and Engineering James A. Newell, 2008-12-15 This book begins with four fundamental tenants The properties of a material are determined by its structure Processing can alter that structure in specific and predictable ways The behavior of materials is grounded in science and is understandable The properties of all materials change over time with use and exposure to environmental conditions When selecting a material sufficient and appropriate testing must be performed to insure that the material will remain suitable throughout the reasonable life of the product This text assumes that the students are at least sophomores so that they are familiar with basic chemical bonding and the periodic table But it is an introductory materials course so there will be no differential equations percolation theory quantum mechanics statistical thermodynamics or other advanced topics The book is designed as an introduction to the field not a comprehensive guide to all materials science knowledge Instead of going into great detail in many areas the book provides key concepts and fundamentals students need to understand materials science and make informed decisions An example of the philosophy is found in the materials testing section Although countless variations exist in testing techniques the chapter focuses on operating principles and the property to be measured rather than confusing the student with exposition on variations and exceptions That material is beyond the scope of most introductory courses

*Books in Print Supplement*, 2002      **Children's Books in Print** R R Bowker Publishing, Bowker, 1999-12      *General Chemistry* Ralph H. Petrucci, 1989      Quantum Mechanics Claude Cohen-Tannoudji, Bernard Diu, Franck Laloë, 2020      Physical Chemistry Thomas Engel, 2006      **Analytical Chemistry** Gary D. Christian, 1980 Extensively revised and updated this edition is concerned primarily with quantitative analysis techniques Describes how to design an analytical method how to obtain a laboratory sample that is representative of the whole and to prepare it for analysis what measurement tools are available automated analyses and the statistical significance of the analysis New and expanded topics include heterogeneous equilibria diode array spectrometers fiber optic sensors and solid phase extraction      **Choice**, 2006      *Secondary Teaching Methods* Kenneth T. Henson, 1981      **Resources in Education**, 1968      *Educational Times and Journal of the College of Preceptors*, 1910      *The Educational Times, and Journal of the College of Preceptors*, 1910      *Education Outlook*, 1910      Research in Education, 1968-07      Soviet Electrochemistry, 1975      *Survival Handbook for the New Chemistry Instructor* Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

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