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CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

According to the law of conservation of mass, the total mass of the reacting substances is

- (a) always more than the total mass of the products.
- (b) always less than the total mass of the products.
- (c) sometimes more and sometimes less than the total mass of the products.
- (d) always equal to the total mass of the products.

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

Dute:

each question.		
2. In the reaction described by the word of	quation	
sodium + w	ater sodium h	ydroxide + hydrogen gas
 a. sodium hydroxide is a product. 	lb.	sodium is a product.
c. water is a reactant.	d.	Both (a) and (c)
Which of the following indicates that a	chemical equati-	on is balanced?
 The numbers of atoms of each elem 	cent are the same	on both sides of the equation.
 All of the coefficients are the same. 		
 c. The numbers of molecules on each 	side of the equat	ion are equal.
d. The sums of the coefficients on each		
 4. From a complete and correctly written 		en, you can obtain the
 a. chemical formulas of the reactants a 		
 relative amounts of the reactants an 		
 e. physical states of the reactants and p 	products.	
d. All of the above		
5. To balance a chemical equation, you m	HIST	
 a. adjust the subscripts. 	lb.	adjust the number of products.
 e. adjust the number of reactants. 	4.	adjust the coefficients.
6. The formation of a solid product in a cl		
a. (aq). b. (g).	c. (/).	d. (r).
8. In the expression 3CO ₂ , the numbers J	ama z are, respec	
 a. a subscript and a coefficient. c. two subscripts. 	d.	a coefficient and a subscript, two coefficients.
		d by the following chemical equation? (Add up the coefficients.)
9. What is the total number of indecedies	NH ₄ NO ₂ -+ 2	
a. 3 b. 4	c. 8	d. 16
10. Which equation below violates the law	of conservation	of mass?
a. $2H_2 + O_2 \rightarrow 2H_2O$	to.	$KCl + Br \rightarrow KBr + Cl_l$
c. $2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$	d.	$Na_{2}CO_{1} + 2HCI \rightarrow 2NaCI + H_{2}O + CO_{2}$
11. Which of the following equations repre	ments the ballance	ed equation for the reaction of iron and oxygen?
a. $2Fe + O_7 \rightarrow Fe_2O_3$	b.	$Fe + 3O_2 \rightarrow Fe_2O_3$
c. $4Fe + 3O_2 \rightarrow 2Fe_2O_3$	d.	$3Fe + 3O_2 \rightarrow 2Fe_2O_2$
12. Which coefficients correctly balance th	e formula equati	on below?
	$KCIO_2(r) \longrightarrow KC$	$C(x) + O_2(y)$
a, 1, 1, 1 b, 1, 1, 3	c. 2, 2, 3	d. 2, 1, 1
13. Which equation is not balanced?		
a. $2H_2 + O_2 \rightarrow 2H_2O$	b	$4H_2 + 2O_2 \rightarrow 4H_2O$
c. H ₂ + H ₂ + O ₂ → H ₂ O + H ₂ O	d.	2H; + O; -> H;O
14. In which type of reaction do two or mo	re compounds re	act to form one product?
 a. synthesis (AKA combination) 	ъ.	decomposition
c. combustion	d.	single-replacement
15. Which of the following is true regarding	g single-replaces	
 a. Only one product is formed. 	b.	There is only one reagent.
 a. It requires oxygen. 	d.	One element replaces another in a compound.

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