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CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

According to the law of conservation of mass, the total mass of the reacting substances is

- (a) always more than the total mass of the products.
- (b) always less than the total mass of the products.
- (c) sometimes more and sometimes less than the total mass of the products.
- (d) always equal to the total mass of the products.

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

Dute:

each question.		
 In the reaction described by the word 	equation	
sodium + ·	water sodium h	ydroxide + hydrogen gas
 a. sodium hydroxide is a product. 	b.	sodium is a product.
c. water is a reactant.	d.	Both (a) and (c)
3. Which of the following indicates that	a chemical equation	on is balanced?
a. The numbers of atoms of each ele-	ment are the same	on both sides of the equation.
 All of the coefficients are the same 	e.	
c. The numbers of molecules on each	h side of the equati	ion are equal.
d. The sums of the coefficients on ea	sch side of the equi	ation are equal.
 From a complete and correctly writter 	n chemical equatio	en, you can obtain the
 a. chemical formulas of the reactants 		
 b. relative amounts of the reactants a 		
 physical states of the reactants and 	d products.	
d. All of the above		
5. To balance a chemical equation, you	must	
 a. adjust the subscripts. 	lb.	adjust the number of products.
 c. adjust the number of reactants. 	d.	adjust the coefficients.
6. The formation of a solid product in a		
a. (aq). b. (g).	e. (/).	d. (s).
8. In the expression 3CO ₂ , the numbers .	3 and 2 are, respec-	
 a. a subscript and a coefficient. 	D	a coefficient and a subscript.
c. two subscripts.	d.	two coefficients.
9. What is the total number of molecules	s that is represente NHaNO ₂ > N	d by the following chemical equation? (Add up the coefficients.)
a. 3 b. 4	c. 8	d. 16
10. Which equation below violates the lav		
a. 2H₂ + O₂ → 2H₂O	D.	$KCI + Br \rightarrow KBr + Cb$
c. 2Fe ₂ O ₂ + 3C → 4Fe + 3CO ₂	d.	$Na_2CO_1 + 2HC1 \rightarrow 2NaC1 + H_2O + CO_2$
		ed equation for the reaction of iron and oxygen?
a. 2Fe + O ₂ → Fe ₂ O ₁	ь.	$Fe + 3O_2 \rightarrow Fe_2O_3$
e. 4Fe + 3O ₂ → 2Fe ₂ O ₁	-4	$3Fe + 3O_2 \rightarrow 2Fe_2O_2$
12. Which coefficients correctly balance t	the formula equati-	
	$KClO_2(x) \rightarrow KC$	
a, 1, 1, 1 b, 1, 1, 3	c. 2, 2, 3	d. 2, 1, 1
13. Which equation is not balanced?		
a. $2H_2 + O_2 \rightarrow 2H_2O$	b.	$4H_2 + 2O_2 \rightarrow 4H_2O$
c. H ₂ + H ₂ + O ₂ → H ₂ O + H ₂ O	at.	$2H_2 + O_2 \rightarrow H_2O$
14. In which type of reaction do two or m	sore compounds re	
a. synthesis (AKA combination)	ъ.	decomposition
c. combustion	d.	single-replacement
15. Which of the following is true regards	ing single-replacer	
 a. Only one product is formed. 	ъ.	There is only one reagent.
 a. It requires oxygen. 	d.	One element replaces another in a compound.

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