

CHEMISTRY 10TH – Solved Numerical

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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1. For the reaction of sulphur dioxide and oxygen to form sulphur trioxide, the balanced reversible reaction is:



According to the law of mass action:

$$\text{Rate of forward reaction: } R_f = k_f [\text{SO}_2]^2 [\text{O}_2]$$

$$\text{Rate of reverse reaction: } R_r = k_r [\text{SO}_3]^2$$

The expression for equilibrium constant k_c is:

$$k_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$$

2. Equilibrium occurs when nitrogen monoxide gas reacts with oxygen gas to form nitrogen dioxide gas.



At equilibrium at 230°C, the concentrations are measured to be:

$$[\text{NO}] = 0.0542 \text{ mol.dm}^{-3} \quad [\text{O}_2] = 0.127 \text{ mol.dm}^{-3} \quad [\text{NO}_2] = 15.5 \text{ mol.dm}^{-3}$$

Calculate the equilibrium constant at this temperature.

Solution:

Given equilibrium concentrations of reactants and product are:

$$[\text{NO}] = 0.0542 \text{ mol.dm}^{-3}$$

$$[\text{O}_2] = 0.127 \text{ mol.dm}^{-3}$$

$$[\text{NO}_2] = 15.5 \text{ mol.dm}^{-3}$$

Write equilibrium expression as:

$$k_c = \frac{[\text{NO}_2]^2}{[\text{NO}]^2 [\text{O}_2]}$$

Now put equilibrium concentration values in equilibrium constant expression

$$k_c = \frac{[15.5 \text{ mol.dm}^{-3}]^2}{[0.0542 \text{ mol.dm}^{-3}]^2 [0.127 \text{ mol.dm}^{-3}]}$$

$$\boxed{K_c = 6.44 \times 10^5 \text{ mol}^{-1} \text{ dm}^3} \quad \text{Ans.}$$

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