

# AP Chemistry Pacing Guide

<b>Unit 4: Chemical Reactions</b> <ul style="list-style-type: none"> <li>Introduction for Reactions</li> <li>Net Ionic Equations</li> <li>Representations of Reactions</li> <li>Physical and Chemical Changes</li> <li>Stoichiometry</li> <li>Introduction to Titration</li> <li>Types of Chemical Reactions</li> <li>Introduction to Acid-Base Reactions</li> <li>Oxidation-Reduction (Redox) Reactions</li> </ul>	<b>BIG IDEA 1:</b> Scale, Proportion, and Quantity <b>BIG IDEA 3:</b> Transformations	<ol style="list-style-type: none"> <li>I can identify evidence of chemical and physical changes in matter.</li> <li>I can represent changes in matter with a balanced chemical or net ionic equation: a. For physical changes; b. For given information about the identity of the reactants and/or product.</li> <li>For ions in a given chemical reaction I can represent a given chemical reaction or physical process with a consistent particulate model.</li> <li>I can explain the relationship between macroscopic characteristics and bond interactions for: a. Chemical processes; b. Physical processes.</li> <li>I can explain changes in the amounts of reactants and products based on the balanced reaction equation for a chemical process.</li> <li>I can identify the equivalence point in a titration based on the amounts of the titrant and analyte, assuming the titration reaction goes to completion.</li> <li>I can identify a reaction as acid-base, oxidation-reduction, or precipitation.</li> <li>I can identify species as Brønsted-Lowry acids, bases, and/or conjugate acid-base pairs, based on proton-transfer involving those species.</li> <li>I can represent a balanced redox reaction equation using half-reactions.</li> </ol>	10 days
<b>Unit 5: Kinetics</b> <ul style="list-style-type: none"> <li>Reaction Rates</li> <li>Introduction to Rate Law</li> <li>Concentration Changes Over Time</li> <li>Elementary Reactions</li> <li>Collision Model</li> <li>Reaction Energy Profile</li> <li>Introduction to Reaction Mechanism</li> <li>Mechanism Reaction Mechanism and Rate Law</li> <li>Steady-State</li> </ul>	<b>BIG IDEA 3:</b> Transformations <b>BIG IDEA 6:</b> Energy	<ol style="list-style-type: none"> <li>I can explain the relationship between the rate of a chemical reaction and experimental parameters.</li> <li>I can represent experimental data with a consistent rate law expression.</li> <li>I can identify the rate law expression of a chemical reaction using data that show how the concentrations of reaction species change over time.</li> <li>I can represent an elementary reaction as a rate law expression using stoichiometry.</li> <li>I can explain the relationship between the rate of an elementary reaction and the frequency, energy, and orientation of molecular collisions.</li> <li>I can represent the activation energy and overall energy change in an elementary reaction using a reaction energy profile.</li> <li>I can identify the components of a reaction mechanism.</li> <li>I can identify the rate law for a reaction from a mechanism in which the first step is rate limiting.</li> <li>I can identify the rate law for a reaction from a mechanism in which the</li> </ol>	10 days

# Modern Chemistry Pacing Guide

**M Mark**



## **Modern Chemistry Pacing Guide:**

## Reviewing **Modern Chemistry Pacing Guide**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Modern Chemistry Pacing Guide**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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