CHAPTER 6 REVIEW

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

1. _A_ The notation for sodium chloride, NaCl, stands for one
(a)formula unit. (c) crystal.
(b)molecule.(d) atom.

2. _D_ In a crystal of an ionic compound, each cation is surrounded by a number of
(a)molecules. (c) dipoles.
(b)positive ions. (d) negative ions.

3. _B_ Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
(a)higher in potential energy. (c) equal in potential energy.
(b)lower in potential energy. (d) unstable.

4. _B_ The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
(a)Compound A is not an ionic compound.
(b)It will be more difficult to break the bonds in compound A than those in compound B.
(c)Compound B has larger crystals than compound A.
(d)Compound A has larger crystals than compound B.

5. _B_ The forces of attraction between molecules in a molecular compound are generally
(a)stronger than the attractive forces among formula units in ionic bonding.
(b)weaker than the attractive forces among formula units in ionic bonding.
(c)approximately equal to the attractive forces among formula units in ionic bonding.
(d)equal to zero.

6.Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

7.What type of energy best represents the strength of an ionic bond?

Lattice energy

8.What types of bonds are present in an ionic compound that contains a polyatomic ion?

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