

## Chapter 9: Stoichiometry

### I. Stoichiometry

A. Relating the mass of the reactants with the mass of the products in a chemical reaction is Stoichiometry.

1. Stoichiometry follows the Law of Conservation of mass.
2. The total sum of the mass for reactants will equal the total sum of the mass of the products and NOT the number of molecules.
3. Stoichiometry is dependent on balanced equations.

B. Stoichiometric questions investigate the amount of substance needed to react with another substance, or the amount of a particular substance produced in a chemical reaction.

Stoichiometric relationships can look at:

1. mole to mole ratios.

a. Mole to mole ratios relate the number of moles of any substance in a balanced chemical reaction. They are based on the balanced chemical equation.



3 moles of  $\text{H}_2$  requires 1 mole of  $\text{N}_2$  to form 2 moles of  $\text{NH}_3$ .

Therefore, 6 moles of  $\text{H}_2$  should equal 4 moles of  $\text{NH}_3$ .

This means that 17.25 moles of  $\text{H}_2$  can produce 11.5 moles of  $\text{NH}_3$ .

$$17.5 \text{ mol H}_2 \left( \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 11.5 \text{ mol NH}_3$$

2. Mole to gram calculations.

a. Using a balanced chemical reaction the moles of one substance is given, and another substance has an equal amount of grams.

b. The equation MUST be balanced!



Using the same equation, how many grams of  $\text{NH}_3$  are produced if 10.57 moles of  $\text{H}_2$  are used in a reaction?

$$10.57 \text{ mol H}_2 \left( \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) \left( \frac{17.04 \text{ g NH}_3}{1 \text{ mol NH}_3} \right) = 120.08 \text{ g NH}_3$$

Mole to Mole ratio  
from the **BALANCED**  
chemical equation!

Mole to gram  
conversion with the  
molar mass from the  
Periodic Table

d. The number of moles does NOT need to be a whole number.

# Modern Chemistry Section Review Stoichiometry Answers

**Norbert T. Porile**



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