

1. Calculate the Molarity of a solution when enough water is added to 35 gr of NaCl to make 444ml of solution.

$$\frac{35 \text{ g NaCl}}{58.5 \text{ g}} \left| \frac{1 \text{ mol NaCl}}{58.5 \text{ g}} \right| = .6 \text{ mol NaCl}$$

$$M = \frac{.6}{.444} = 1.3 \text{ M NaCl solution}$$

2. A 1.5M solution of sodium chloride was made with 200ml of water. How many moles of salt were used in making the solution?

$$1.5 \text{ M} = \frac{n}{.200 \text{ L}}$$

$$n = .3 \text{ mol NaCl}$$

3. How many grams of salt are there in 200ml of a 1.5M NaCl solution?

$$\frac{.3 \text{ mol NaCl}}{1 \text{ mol}} \left| \frac{58.5 \text{ g NaCl}}{1 \text{ mol}} \right| = 17.55 \text{ g NaCl}$$

4. 120ml of a 2M NaCl solution is added to 120ml of a 1.2M NaCl solution. What is the new Molarity?

$$2 \text{ M} = \frac{n}{.120 \text{ L}} \quad \left[n_1 = .24 \text{ mol} \right]$$

$$1.2 \text{ M} = \frac{n}{.120 \text{ L}} \quad \left[n_2 = .144 \text{ mol} \right]$$

$$n_{\text{total}} = .384 \text{ mol}$$

$$V_{\text{total}} = .240 \text{ L}$$

$$M = \frac{.384 \text{ mol}}{.240 \text{ L}} = 1.6 \text{ M NaCl}$$

5. 400ml of water is added to one liter of a 3M NaCl solution. What is the new Molarity?

$$M_1 V_1 = M_2 V_2$$

$$(3 \text{ M})(1 \text{ L}) = M_2(1.4 \text{ L})$$

$$M_2 = 2.1 \text{ M NaCl}$$

6. How many ml of a 3M NaCl solution are needed to make 222ml of a 1.7M solution?

$$M_1 V_1 = M_2 V_2$$

$$(3 \text{ M}) V_1 = (1.7 \text{ M})(222 \text{ mL})$$

$$\boxed{V_1 = 125.8 \text{ mL}}$$

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Table of Contents Molarity Pogil Packet Answers

1. Understanding the eBook Molarity Pogil Packet Answers
 - The Rise of Digital Reading Molarity Pogil Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Pogil Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Pogil Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Pogil Packet Answers
 - Personalized Recommendations
 - Molarity Pogil Packet Answers User Reviews and Ratings
 - Molarity Pogil Packet Answers and Bestseller Lists
5. Accessing Molarity Pogil Packet Answers Free and Paid eBooks
 - Molarity Pogil Packet Answers Public Domain eBooks
 - Molarity Pogil Packet Answers eBook Subscription Services
 - Molarity Pogil Packet Answers Budget-Friendly Options
6. Navigating Molarity Pogil Packet Answers eBook Formats

- ePub, PDF, MOBI, and More
- Molarity Pogil Packet Answers Compatibility with Devices
- Molarity Pogil Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Pogil Packet Answers
 - Highlighting and Note-Taking Molarity Pogil Packet Answers
 - Interactive Elements Molarity Pogil Packet Answers
- 8. Staying Engaged with Molarity Pogil Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Pogil Packet Answers
- 9. Balancing eBooks and Physical Books Molarity Pogil Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Pogil Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Molarity Pogil Packet Answers
 - Setting Reading Goals Molarity Pogil Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Molarity Pogil Packet Answers
 - Fact-Checking eBook Content of Molarity Pogil Packet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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