

# MOLARITY BY DILUTION

79 + 52 = 131 Name \_\_\_\_\_

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid =  $M \times V$  then,  
 $M_1 \times V_1 = M_2 \times V_2$  Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(250 mL)}{18 M}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 mL)}{12 M}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 mL)}{3.0 M}$$

125 mL

4. To how much water should 50 mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 mL)}{4.0 M} = 150 mL - 50 mL$$

100 mL H<sub>2</sub>O

5. To how much water should 100 mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 mL)}{1.5 M} = 1200 mL - 100 mL = 1100 mL$$

1.1 L

# Molarity Answer Key Chemistry If8766

**RJ Alexander**



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## **Table of Contents Molarity Answer Key Chemistry If8766**

1. Understanding the eBook Molarity Answer Key Chemistry If8766
  - The Rise of Digital Reading Molarity Answer Key Chemistry If8766
  - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Answer Key Chemistry If8766
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Molarity Answer Key Chemistry If8766
  - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Answer Key Chemistry If8766
  - Personalized Recommendations
  - Molarity Answer Key Chemistry If8766 User Reviews and Ratings
  - Molarity Answer Key Chemistry If8766 and Bestseller Lists
5. Accessing Molarity Answer Key Chemistry If8766 Free and Paid eBooks
  - Molarity Answer Key Chemistry If8766 Public Domain eBooks
  - Molarity Answer Key Chemistry If8766 eBook Subscription Services
  - Molarity Answer Key Chemistry If8766 Budget-Friendly Options

6. Navigating Molarity Answer Key Chemistry If8766 eBook Formats
  - ePub, PDF, MOBI, and More
  - Molarity Answer Key Chemistry If8766 Compatibility with Devices
  - Molarity Answer Key Chemistry If8766 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Molarity Answer Key Chemistry If8766
  - Highlighting and Note-Taking Molarity Answer Key Chemistry If8766
  - Interactive Elements Molarity Answer Key Chemistry If8766
8. Staying Engaged with Molarity Answer Key Chemistry If8766
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Molarity Answer Key Chemistry If8766
9. Balancing eBooks and Physical Books Molarity Answer Key Chemistry If8766
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Molarity Answer Key Chemistry If8766
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Molarity Answer Key Chemistry If8766
  - Setting Reading Goals Molarity Answer Key Chemistry If8766
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Answer Key Chemistry If8766
  - Fact-Checking eBook Content of Molarity Answer Key Chemistry If8766
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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