

MOLARITY BY DILUTION

79 + 52 = 131 Name _____

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = $M \times V$ then,
 $M_1 \times V_1 = M_2 \times V_2$ Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 M)(250 mL)}{18 M}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 M)(100 mL)}{12 M}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 M)(25 mL)}{3.0 M}$$

125 mL

4. To how much water should 50 mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 M)(50 mL)}{4.0 M} = 150 mL - 50 mL$$

100 mL H₂O

5. To how much water should 100 mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 M)(100 mL)}{1.5 M} = 1200 mL - 100 mL = 1100 mL$$

1.1 L

Molarity By Dilution Answer Key

Lorraine J. Doucette

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Molarity By Dilution Answer Key:

General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition. Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value.

Mathematics for the Clinical Laboratory E-Book Lorraine J. Doucette, 2020-06-18 Selected for Doody's Core Titles 2024 with Essential Purchase designation in Laboratory Technology. Master the skills you'll need to perform accurate clinical laboratory calculations. Mathematics for the Clinical Laboratory 4th Edition demonstrates the calculations used in the analysis of test specimens. It begins by explaining basic mathematical principles and then covers the types of calculations needed in specific areas of the clinical lab including urinalysis, hematology and microbiology. Finally it focuses on the statistical calculations used in quality assurance and quality control. Step by step examples reinforce your understanding and calculation templates and practice problems ensure that you make correct calculations every time. Step by step examples explain basic mathematical principles and show you exactly how to perform each type of calculation. Sample problems with answers can also be used as templates for solving laboratory calculations. Practice problems at the end of each chapter provide a self assessment tool helping you determine what you need to review. Summaries of important formulas are included at the end of the text's major sections. Coverage of statistical calculations includes standard deviation as well as calculations associated with quality assurance and quality control. Quick tips and notes make it easier to understand and remember pertinent information. Learning objectives at the beginning of each chapter provide measurable outcomes to achieve by completing the chapter material. Full color design includes 100 illustrations. Useful appendix of Greek symbols provides a quick reference to turn to when studying. Glossary at the back of the textbook includes definitions of important mathematical terms. New Updated content and calculations reflect the latest procedures used in today's laboratories.

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Chemistry John Christian Bilar, 1989

Basic Chemistry Steven S. Zumdahl, 1999 For a full description see catalog entry for Zumdahl Introductory Chemistry A Foundation 4 e

Mathematics for the Clinical Laboratory Lorraine J. Doucette, 1997 This comprehensive practical text shows you how to perform the clinical calculations used in each area of the laboratory to ensure accurate results for patients

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Keys to Engineering Success Jill S. Tietjen, 2001 Lively in format and filled with real world vignettes applications and examples this introduction to engineering is designed to keep engineering students encouraged and motivated during their freshmen year when they can't yet see how all of the calculus physics and chemistry relates to their later education and careers as engineers The real world vignettes and pictures capture not only the diversity of the profession but of the engineers themselves providing an overview of the various types of engineering as well as what working professionals do The book also features extensive information on engineering specific study skills gives hints and suggestions on how to enhance one's college experience and provides information on what resources to look for and where to find them Includes extensive skill building exercises on perforated pages So What Is Engineering Anyway Where to Get Help When You Need It Critical and Creative Thinking Tapping the Power of Your Mind Reading and Study Skills Writing Listening and Test Taking Goal Setting and Time Management Relating to Others Appreciating Your Diverse World Managing Career and Money Reality Resources Preparing for the Changes in the Engineering Field For students just entering an engineering program

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Iterations, II Russell Batt, John W. Moore, 1987

Chemistry James E. Brady, John R. Holum, 1993-01-21 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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chart that shows the classification of elements is called the. Properties of Atoms and the Periodic Table 37. Assessment. Page 6. Assessment. Name. Chapter ...