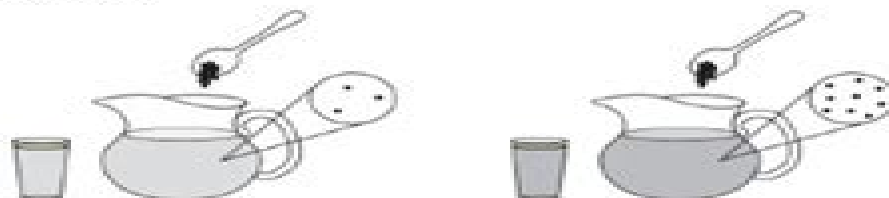


When you buy a bottle of a certain brand of lemonade you expect it to taste just as sweet as the last time you bought that kind of lemonade. Likewise, when doctors prescribe a certain ointment, they expect the concentration of medicine to be consistent. How do companies ensure that their products taste or perform the same every time you purchase them? Many companies, including pharmaceutical companies, keep track of the concentration of a solution by measuring its **molarity** – a ratio of the number of solute particles to the volume of the solution. In this activity you will learn about molarity and how to represent concentration quantitatively.

Model 1 – Lemonade Mixtures*



* Both pitchers were filled with enough water (solvent) to provide 2 L of solution. The solid lemonade mixture consists of several molecules. The dissolved sugar molecule (solute) is indicated with a •

1. Refer to Model 1.
 - a. What is the solvent in this scenario? water The solute? sugar
 - b. Lemonade Solution 1 has (more/less/the same) volume of solution as Solution 2.
 - c. Lemonade Solution 1 has (more/less/the same) quantity of solute as Solution 2.
2. Lemonade Solution 2 is considered to be **concentrated**, and Lemonade Solution 1 is considered to be **dilute**. Examine the two pictures in Model 1. List two ways to differentiate a concentrated solution from a dilute solution.

Visually; the more concentrated solution will have a deeper color. Taste! The more concentrated solution will have a sweeter (in this case) taste.

Molarity Pogil Activities For High School Chemistry

CL Gary



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