

Name: Mia Garcia Date: _____ Period: _____

What happened to
Avogadro when he
got bit by 6.02×10^{23} mosquitoes?

He got Mol-aria

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry**- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Answer:

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Mole ratio

OR

$$34.8 \text{ mol NH}_3 \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Mole ratio

Key Questions:

1. What was the given in the example above?

34.8 mol of NH_3

2. Where did the $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?

Mole Ratio

3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.

Go the other
direction
to check

$$\frac{52.2 \times 2}{3} = 34.8$$

Molarity Pogil Answers

Richard Samuel Moog



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Process Oriented Guided Inquiry Learning (POGIL) Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL s theoretical basis its implementations in diverse environments and evaluation of student outcomes

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