

Molarity Practice

Calculate for the unknown variables. $M = n/v$

1. Explain what molarity/concentration means?
2. What is the molarity of a solution that contains 5.50 moles of calcium chloride in 2.5 liters of solution?
3. What is the molarity of a solution that contains 3.39×10^{-2} moles hydrobromic acid in 1 liter of solution?
4. A beaker contains 250g $C_6H_{12}O_6$ (glucose) in 1,500 mL of solution. What is the molarity?
5. A flask contains 325g grams iron II oxide and has a molarity of 2M. What is the volume of the solution?
6. Calculate the moles if a flask contains 3.25M potassium nitrate in 750 mL of solution.
7. A cup contains 85.31 grams ammonium acetate in 3.27 liters of solution - calculate the molarity.
8. What is the molarity of a solution that contains 10.5 grams magnesium oxide in 250 mL of solution?
9. What mass of potassium phosphate would you use to make 2.5 liter of 1.25 M solution?

Molarity Practice Unit Lesson 2

Christian G. Meyer



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