Molar Mass and Mole Calculations Worksheet

- Calculate the molar mass for each of the following compounds:
 - a. NaCl = 58.44 g/mol

e. $(NH_4)_3PO_4 = 149.12 \text{ g/mol}$

b. Fe₂O₃ = 159.7 g/mol

f. Ca₀(PO₄)₂=310.18 g/mol

c. MgCO₃ = 84.32 g/mol

g. $C_6H_{12}O_6 = 180.18$ g/mol

d. Ba(OH)2 = 171.35 g/mol

- h. CsH₅N = 79.11 g/mol
- 2. Calculate the number of moles in each of the following masses:
 - a. 45.0 g of acetic acid, CH₃COOH = 0.75 mol
 - b. 7.04 g of lead (II) nitrate = 0.02 mol
- 3. Calculate the mass in grams of each of the following amounts:
 - a. 3.00 mol of selenium oxybromide, SeOBr2 = 764.31 g

M=78.97+ 16+2(79.9)= 254.77 g/mol 254.77 x 3=

b. 488 mol of calcium carbonate = 48,843.92 g

M=40.08+12.01+3(16)= 100.09 g/mol 100.09 x 488=

4. Calculate the number of molecules or formula units in each of the following

amounts:

a. 4.27 mol of tungsten(VI) oxide, WO₃ = 2.57 x 10²⁴
4.27 x 6.02x10²³=

b. 0.989 mol of nitrobenzene C₆H₅NO₂ = 5.95 x 10²³

- b. 0.989 mol of nitrobenzene $C_6H_5NO_2 = 5.95 \times 10^{-9}$ $0.989 \times 6.02 \times 10^{23} =$
- Calculate the mass of each of the following quantities:
 - a. 8.39×10^{23} molecules of fluorine = 52.96 g

8.39x19²³/6.02x10²³= 1.39 mol 1.39 x 38=

M= 2(19) = 38 g/mol

b. 6.82×10^{24} formula units of beryllium sulfate = 1,190.33 g

M = 9.01 + 32.06 + 4(16) = 105.07 $105.07 \times 11.33 = 6.82 \times 10^{24} / 6.02 \times 10^{23} = 11.33$

Mole And Mass Answer Key

Lei Huang

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