

Mole to Mole Stoichiometry Worksheet

- 1) In the reaction below, how many moles of aluminum will be used when reacted with 1.33 moles of oxygen?
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$

$$1.33 \text{ mol O}_2 \times \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = \boxed{1.77 \text{ mol Al}}$$

- 2) How many moles of hydrogen will be produced when reacted with 0.0204 moles of sodium in the reaction?
 $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + 1 \text{ H}_2$

$$0.0204 \text{ mol Na} \times \frac{1 \text{ mol H}_2}{2 \text{ mol Na}} = \boxed{0.0102 \text{ mol H}_2}$$

- 3) Using the equation $2 \text{ Al(NO}_3)_3 + 3 \text{ MgBr}_2 \rightarrow 3 \text{ Mg(NO}_3)_2 + 2 \text{ AlBr}_3$
 a. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 2.3 moles of $\text{Mg(NO}_3)_2$

$$2.3 \text{ mol Mg(NO}_3)_2 \times \frac{2 \text{ mol Al(NO}_3)_3}{3 \text{ mol Mg(NO}_3)_2} = \boxed{1.53 \text{ mol Al(NO}_3)_3}$$

- b. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 0.78 moles of AlBr_3

$$0.78 \text{ mol AlBr}_3 \times \frac{2 \text{ mol Al(NO}_3)_3}{2 \text{ mol AlBr}_3} = \boxed{0.78 \text{ mol Al(NO}_3)_3}$$

- c. Find the number of moles MgBr_2 that reacts with 1.05 moles of $\text{Al(NO}_3)_3$

$$1.05 \text{ mol Al(NO}_3)_3 \times \frac{3 \text{ mol MgBr}_2}{2 \text{ mol Al(NO}_3)_3} = \boxed{1.58 \text{ mol MgBr}_2}$$

- 4) In the reaction below, how many moles of carbon dioxide are produced if 19.65 moles of water are produced?
 $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$

$$19.65 \text{ moles H}_2\text{O} \times \frac{4 \text{ mol CO}_2}{6 \text{ mol H}_2\text{O}} = \boxed{13.1 \text{ mol CO}_2}$$

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ChemDiscovery Teacher Edition Olga I. Agapova, 2002 Chemistry All-in-One For Dummies (+ Chapter Quizzes Online) Christopher R. Hren, John T. Moore, Peter J. Mikulecky, 2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you'll need to pass your chemistry class with flying colors You'll find tons of practical examples and practice problems and you'll get access to an online quiz for every chapter Reinforce the concepts you learn in the classroom and beef up your understanding of all the chemistry topics covered in the standard curriculum Prepping for the AP Chemistry exam Dummies has your back with plenty of review before test day With clear definitions concise explanations and plenty of helpful information on everything from matter and molecules to moles and measurements Chemistry All in One For Dummies is a one stop resource for chemistry students of all valences Review all the topics covered in a full year high school chemistry course or one semester of college chemistry Understand atoms molecules and the periodic table of elements Master chemical equations solutions and states of matter Complete practice problems and end of chapter quizzes online Chemistry All In One For Dummies is perfect for students who need help with coursework or want to cram extra hard to ace that chem test **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity *Foundations of College Chemistry* Morris Hein, Susan Arena, Cary Willard, 2016-08-02 This text is an unbound three hole punched version Used by over 750 000 students Foundations of College Chemistry Binder Ready Version 15th Edition is praised for its accuracy clear no nonsense approach and direct writing style Foundations direct and straightforward explanations focus on problem solving making it the most dependable text on the market Its comprehensive scope proven track record outstanding in text examples and problem sets were all designed to provide instructors with a solid text while not overwhelming students in a difficult course Foundations fits into the prep intro chemistry courses which often include a wide mix of students from science majors not yet ready for general chemistry allied health students in their 1st semester of a GOB sequence science education students for elementary school teachers to the

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