



Stoichiometry Challenge

Written by Dolores Dang-Wright and Sharm Kanner, 2010 Workshop Participants

Teacher Background Information:

This lesson replaces a traditional Aluminum to Alum stoichiometry lab with a greener precipitation reaction of sodium carbonate and calcium chloride. It is used to demonstrate how stoichiometry works, showing that if concentrations and amounts of the starting materials are known that the theoretical yield can be calculated from a balanced chemical equation.

Safety information:

- Calcium chloride (CaCl_2) and sodium carbonate (Na_2CO_3) are both skin irritants.
- If either chemical comes in contact with skin wash with soap and water.
- If they come in contact with your eyes spend 15 minutes in the eyewash and call a doctor.

Educational Goals: Students will understand...

- How to calculate and make molar solutions of CaCl_2 and Na_2CO_3 in order to produce an assigned amount of CaCO_3 .
- How to research where CaCO_3 can be found nature.

Student Objectives: Students will ...

- Synthesize an assigned amount of calcium carbonate.
- Demonstrate the mass-mass relationship in chemical reactions.
- Calculate the percent yield and analyze sources of error.
- Practice lab safety

Materials: (per lab group -3 students)

- | | |
|--------------------------------------|------------------------------|
| • 2.2 g CaCl_2 (s) | • Funnel rack |
| • 2.1 g Na_2CO_3 (s) | • 125-mL Erlenmeyer flask |
| • 40 mL Distilled water | • Wax paper or aluminum foil |
| • Three 50-mL or 100-mL beakers | • Stirring rod |
| • Filter paper | • Spatula |
| • Funnel | • 2 plastic weigh boats |
| | • Balance |

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