



## Unit: Stoichiometry **Answer key** "Mole-Mole Calculations" - WS #1

**Directions:** For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide ( $K_2O$ ) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol } K \times \frac{2 \text{ mol } K_2O}{4 \text{ mol } K} = 0.760 \text{ mol } K_2O$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol } Cu_2S \times \frac{2 \text{ mol } Cu}{1 \text{ mol } Cu_2S} = 0.50 \text{ mol } Cu$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol } Na_2SO_4 \times \frac{2 \text{ mol } Na}{1 \text{ mol } Na_2SO_4} = 7.50 \text{ mol } Na$$

3. How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol } Fe_2O_3 \times \frac{3 \text{ mol } CaO}{1 \text{ mol } Fe_2O_3} = 4.8 \text{ mol } CaO$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: **magnesium chlorate  $\rightarrow$  magnesium chloride + oxygen**

Formula equation:  **$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$**   
(must be balanced!)

$$5.12 \text{ mol } Mg(ClO_3)_2 \times \frac{3 \text{ mol } O_2}{1 \text{ mol } Mg(ClO_3)_2} = 15.4 \text{ mol } O_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: **chromium + lead (IV) nitrate  $\rightarrow$  chromium (III) nitrate + lead**

Formula equation:  **$4Cr + 3Pb(NO_3)_4 \rightarrow 4Cr(NO_3)_3 + 3Pb$**   
(must be balanced!)

$$0.85 \text{ mol } Pb(NO_3)_4 \times \frac{4 \text{ mol } Cr(NO_3)_3}{3 \text{ mol } Pb(NO_3)_4} = 1.1 \text{ mol } Cr(NO_3)_3$$

# Lesson 28 Stoichiometry 1 Answer Key

**Robert K. Wismer**



## Lesson 28 Stoichiometry 1 Answer Key:

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *DAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests* Kaplan Test Prep, 2016-10-04 2 full length online practice tests Cover **OAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests** Kaplan Test Prep, 2016-10-04 Kaplan's OAT 2017 2018 Strategies Practice Review provides the content review test taking strategies and realistic practice you need to get the OAT results you want Updated for the latest test changes OAT 2017 2018 is your guide to facing Test Day with confidence The Best Review Two full length online practice tests More than 600 practice questions for every subject with detailed answers and explanations 16 page tear out full color study sheets for quick review on the go A guide to the current OAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the OAT Biology General Chemistry Organic Chemistry Reading Comprehension Physics Quantitative Reasoning Kaplan's proven strategies for Test Day success Expert Guidance Kaplan's expert psychometricians ensure our practice questions and study materials are true to the test We invented test prep Kaplan www.kaptest.com has been helping students for almost 80 years Our proven strategies have helped legions of students achieve their dreams **DAT Prep Plus 2023-2024** Kaplan Test Prep, 2023-02-07 Kaplan's DAT Prep Plus 2023 2024 provides the test taking strategies realistic practice and expert guidance you need to score higher on the Dental Admissions Test Our comprehensive subject review reflects recent changes to the blueprint of the exam question types and test interface You'll get two full length practice DATs and expert tips to help you face Test Day with confidence We're so confident that DAT Prep Plus offers all the knowledge you need to excel at the DAT that we guarantee it after studying with our online resources and book you'll score higher on the DAT or you'll get your money back The Best Review Two updated full length online practice exams for test like practice Study planning guidance More than 600 practice questions for every subject with detailed answers and explanations 12 page full color study sheets for high yield review on the go A guide to the current DAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the DAT Expert Guidance Our books and practice questions are written by veteran teachers who know students every explanation is written to help you learn Kaplan's experts ensure our practice questions and study materials are true to the test We invented test prep Kaplan www.kaptest.com has been helping students for 80 years and our proven strategies have helped legions of students achieve their dreams Publisher's Note Products purchased from 3rd party sellers are not guaranteed by the

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Solutions Guide for Chemistry Steven S. Zumdahl, Kenneth C. Brooks, 1989 *Student Study Guide to Accompany Petrucci's General Chemistry* Robert K. Wismer, 1985

**An Introduction to General, Organic, and Biological Chemistry** Norman L. Allinger, Melvin Jerome Bigelow, Harmon C. McAllister, 1976 **Introductory Chemistry** Charles H. Corwin, 2005 For one semester courses in Basic Chemistry, Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry. This text is carefully crafted to help students learn chemical skills and concepts more effectively. Corwin covers math and problem solving early in the text; he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs.

**Advanced Placement Examination in Chemistry** Frederick J. Rowe, 1990 New for 1990, this is the only preparation guide available for this exam. Includes analyses of past tests, review of chemistry topics and samples of all multiple choice and free response problems.

**Problems for Introductory University Chemistry** James Newton Butler, Basil A. Dunell, Lionel G. Harrison, 1967 *Report*, 1971 Analytical Chemistry Donald J. Pietrzyk, Clyde Frank, 1979-02-22 Good. No Highlights. No Markup. All pages are intact. Slight Shelfwear. May have the corners slightly dented. May have slight color changes. Slightly damaged spine.

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