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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	413

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of one N-N bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4H_2) + 1(N_2)$$

$$4(436) + 1(945)$$

$$3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = -92$$

$$3150 - x = -92$$

$$x = 3150 + 92$$

$$x = 3242$$
- Calculate the energy of one N-N bond.

$$6 \times \text{N-H} = 2478$$

$$\text{N-N} = \frac{3242 - 2478}{1}$$

$$764 \text{ kJ/mol}$$

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Table of Contents Marking Scheme For 0620 31 M J 14

1. Understanding the eBook Marking Scheme For 0620 31 M J 14
 - The Rise of Digital Reading Marking Scheme For 0620 31 M J 14
 - Advantages of eBooks Over Traditional Books
2. Identifying Marking Scheme For 0620 31 M J 14
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Marking Scheme For 0620 31 M J 14
 - User-Friendly Interface
4. Exploring eBook Recommendations from Marking Scheme For 0620 31 M J 14
 - Personalized Recommendations
 - Marking Scheme For 0620 31 M J 14 User Reviews and Ratings
 - Marking Scheme For 0620 31 M J 14 and Bestseller Lists
5. Accessing Marking Scheme For 0620 31 M J 14 Free and Paid eBooks
 - Marking Scheme For 0620 31 M J 14 Public Domain eBooks

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- Marking Scheme For 0620 31 M J 14 Budget-Friendly Options
- 6. Navigating Marking Scheme For 0620 31 M J 14 eBook Formats
 - ePub, PDF, MOBI, and More
 - Marking Scheme For 0620 31 M J 14 Compatibility with Devices
 - Marking Scheme For 0620 31 M J 14 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Marking Scheme For 0620 31 M J 14
 - Highlighting and Note-Taking Marking Scheme For 0620 31 M J 14
 - Interactive Elements Marking Scheme For 0620 31 M J 14
- 8. Staying Engaged with Marking Scheme For 0620 31 M J 14
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Marking Scheme For 0620 31 M J 14
- 9. Balancing eBooks and Physical Books Marking Scheme For 0620 31 M J 14
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Marking Scheme For 0620 31 M J 14
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Marking Scheme For 0620 31 M J 14
 - Setting Reading Goals Marking Scheme For 0620 31 M J 14
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Marking Scheme For 0620 31 M J 14
 - Fact-Checking eBook Content of Marking Scheme For 0620 31 M J 14
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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