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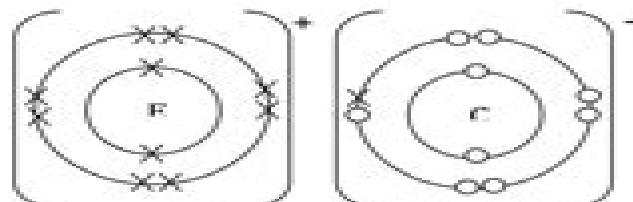
CHEMISTRY PAPER 2 (233/2)

FORM FOUR (4)

TIME: 2 HOURS

MARKING SCHEME

1. 1 (a) Alkaline-earth metals (1mk)
 (b) B (1mk)
 (c) (i) A_2B_3 (½mk) (ii) H_2SO_4 (½mk)
 (d)



Distribution of electrons (1mk) and charges (1mk)

No label (0mk)

- (e) D has a smaller atomic size than C // C has larger atomic size than D (1mk)
 D has more protons than C // D has stronger nuclear charge than C (1mk)
 (f) H (½mk). Loses electrons most readily / largest atomic radius / weakest nuclear charge (explanation ½mk)
 (g) $2A_{(s)} + 6HCl_{(aq)} \longrightarrow 2AlCl_{3(aq)} + 3H_{2(g)}$ 1mk
 Moles of A = $\frac{0.081}{27} = 0.003$ moles ½mk
 Mole ratio 1:3
 Moles of HCl = $0.003 \times 3 = 0.009$ moles ½mk
 $\frac{0.009 \times 1000}{20}$ ½mk = 0.45M ½mk
 (h) The melting point increase with increase in atomic number. 1mk

This is due to increase in the strength of van der waal's forces (inter-molecular forces) 1mk

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