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(Chapter 11) (Alcohols Phenols and Ethers) (Exercises)

(Class 12)



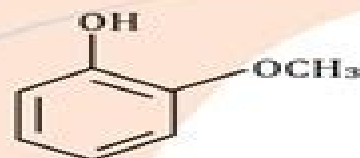
Question 11.15:

Explain why is *ortho* nitrophenol more acidic than *ortho* methoxyphenol?

Answer 11.15:



o - Nitrophenol



o - Methoxyphenol

The nitro-group is an electron-withdrawing group. The presence of this group in the *ortho* position decreases the electron density in the O-H bond. As a result, it is easier to lose a proton. Also, the *o*-nitrophenoxide ion formed after the loss of protons is stabilized by resonance. Hence, *ortho* nitrophenol is a stronger acid.

On the other hand, methoxy group is an electron-releasing group. Thus, it increases the electron density in the O-H bond and hence, the proton cannot be given out easily.

For this reason, *ortho*-nitrophenol is more acidic than *ortho*-methoxyphenol.

Question 11.16:

Explain how does the -OH group attached to a carbon of benzene ring activate it towards electrophilic substitution?

Answer 11.16:

The -OH group is an electron-donating group. Thus, it increases the electron density in the benzene ring as shown in the given resonance structure of phenol.



As a result, the benzene ring is activated towards electrophilic substitution.

Question 11.17:

Give equations of the following reactions:

- (i) Oxidation of propan -1-ol with alkaline KMnO₄ solution.
- (ii) Bromine in CS₂ with phenol.
- (iii) Dilute HNO₃ with phenol.
- (iv) Treating phenol with chloroform in presence of aqueous NaOH.

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**Herbert S. Harned, Benton Brooks
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Table of Contents Mhr Chemistry 12 Solutions

1. Understanding the eBook Mhr Chemistry 12 Solutions
 - The Rise of Digital Reading Mhr Chemistry 12 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Mhr Chemistry 12 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mhr Chemistry 12 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mhr Chemistry 12 Solutions
 - Personalized Recommendations
 - Mhr Chemistry 12 Solutions User Reviews and Ratings
 - Mhr Chemistry 12 Solutions and Bestseller Lists
5. Accessing Mhr Chemistry 12 Solutions Free and Paid eBooks
 - Mhr Chemistry 12 Solutions Public Domain eBooks

- Mhr Chemistry 12 Solutions eBook Subscription Services
- Mhr Chemistry 12 Solutions Budget-Friendly Options
- 6. Navigating Mhr Chemistry 12 Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Mhr Chemistry 12 Solutions Compatibility with Devices
 - Mhr Chemistry 12 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mhr Chemistry 12 Solutions
 - Highlighting and Note-Taking Mhr Chemistry 12 Solutions
 - Interactive Elements Mhr Chemistry 12 Solutions
- 8. Staying Engaged with Mhr Chemistry 12 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mhr Chemistry 12 Solutions
- 9. Balancing eBooks and Physical Books Mhr Chemistry 12 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mhr Chemistry 12 Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mhr Chemistry 12 Solutions
 - Setting Reading Goals Mhr Chemistry 12 Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mhr Chemistry 12 Solutions
 - Fact-Checking eBook Content of Mhr Chemistry 12 Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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