
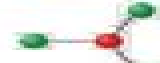
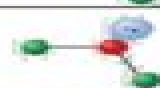







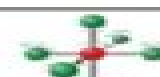




Molecular Geometry

Formula: AB_e A = central atom, B = directly bonded atoms to A, and e = nonbonding (unshared) pairs of electrons

*Note that a molecule formed by joining only **two (2) atoms together** is **linear** regardless of the number of unshared pairs of electrons (AB , ABe , ABe_3 , etc).

AB_e	# of Electron Regions	Electron Geometry	# of Bonding Regions	# of Nonbonding Regions	Molecular Geometry	Structural representation	Hybrid Orbitals	Examples
AB_2	2	Linear	2	0	Linear		sp	$HgCl_2$, CO_2 , HCN
AB_3	3	Trigonal Planar	3	0	Trigonal Planar		sp^2	BF_3 , BCl_3 , SO_3 , CO_3^{2-}
AB_2e	3		2	1	Bent		sp^2	SO_2 , NO_2^-
AB_4	4	Tetrahedral	4	0	Tetrahedral		sp^3	CH_4 , $SiCl_4$, $POCl_3$
AB_3e	4		3	1	Trigonal pyramidal		sp^3	NH_3 , PF_3
AB_2e_2	4		2	2	Bent		sp^3	H_2O , H_2S , BrO_2^-
AB_5	5	Trigonal Bipyramidal	5	0	Trigonal Bipyramidal		dsp^3	PH_5 , PCl_5 , SbF_5 , $IO_3F_2^-$
AB_4e	5		4	1	Distorted tetrahedron ("See-Saw")		dsp^3	SF_6 , IF_4^+
AB_3e_2	5		3	2	T-Shape		dsp^3	ClF_3 , BrF_3
AB_2e_3	5		2	3	Linear		dsp^3	I_2 , ICl_2^- , XeF_2
AB_6	6	Octahedral	6	0	Octahedral		d^2sp^3	SF_6 , PF_6^-
AB_5e	6		5	1	Square Pyramidal		d^2sp^3	IF_5 , $XeOF_4$
AB_4e_2	6		4	2	Square Planar		d^2sp^3	XeF_4 , BrF_4^-

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Vols for 1964 have guides and journal lists Chemical Bonding and Molecular Geometry Ronald James Gillespie, Paul L. A.
Popelier, 2001 Provides an introduction to models and theories of chemical bonding and geometry as applied to the molecules
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hydrogen halogen and methyl derivatives of the elements in each group some of these molecules are ionic some polar
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of molecules that do not conform We also describe electron donor acceptor and hydrogen bonded complexes Chemists use
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been studied The book provides a lucid discussion of a number of models such as the Lewis electron pair bond and the
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relevant to the arrangement of atoms The book is comprised of seven chapters that tackle several areas of molecular
geometry Chapter 1 reviews the definition and determination of molecular geometry while Chapter 2 discusses the unified
view of stereochemistry and stereochemical changes Chapter 3 covers the geometry of molecules of second row atoms and
Chapter 4 deals with the main group elements beyond the second row The book also talks about the complexes of transition
metals and f block elements and then covers the organometallic compounds and transition metal clusters The last chapter
tackles the consequences of small local variations in geometry The text will be of great use to chemists who primarily deal
with the properties of molecules and atoms *A Pictorial Approach to Molecular Bonding* John G. Verkade, 2012-12-06 With
the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are
currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate
the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as
associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of

general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example Lewis Structures and Molecular Geometry (computer File). Jeffrey R. Buell, 1968

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