

# CHEMISTRY AND PHYSICS PRACTICALS

## CHEMISTRY

1. You are provided with ammonium chloride ( $\text{NH}_4\text{Cl}$ ), sodium hydroxide pellets ( $\text{NaOH}$ ), measuring cylinder, thermometer, stirring rod, distilled water, spatula.
  - a) Pour  $50\text{ cm}^3$  of distilled water into one test tube.
  - b) Measure the initial temperature in the test tube and record in the table of results.
  - c) Add half a spatula of ammonium chloride ( $\text{NH}_4\text{Cl}$ ) into the test tube and stir gently.
  - d) Measure the final temperature of the solution in the test tube and record the result.
  - e) Remove the thermometer from the test tube and rinse it with water.
  - f) Repeat the steps (a) to (d) with sodium hydroxide pellets and record the results.

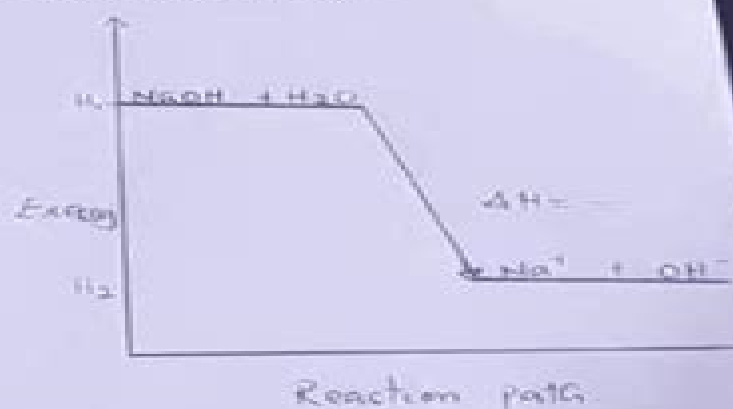
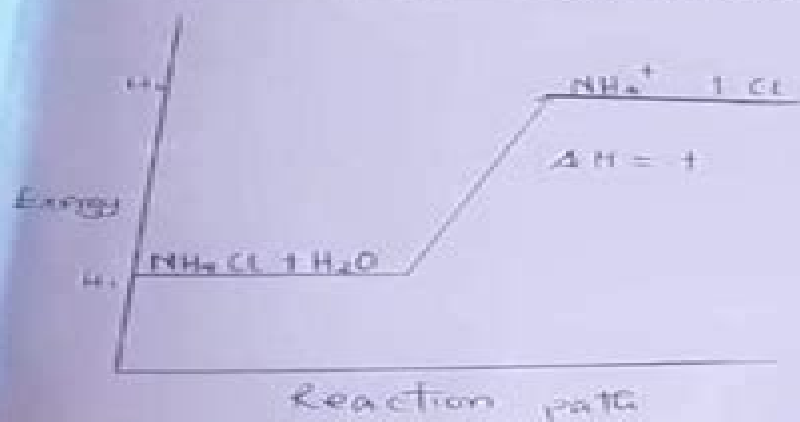
Solution	Initial ( $^{\circ}\text{C}$ ) temperature	Final ( $^{\circ}\text{C}$ ) temperature	Temperature change (final - initial)
$\text{NH}_4\text{Cl}$	$28^{\circ}\text{C}$	$25^{\circ}\text{C}$	$-3^{\circ}\text{C}$
$\text{NaOH}$	$28^{\circ}\text{C}$	$45^{\circ}\text{C}$	$17^{\circ}\text{C}$

- g) Classify the reactions as exothermic or endothermic.

Water + ammonium chloride Endothermic

Water + sodium hydroxide Exothermic

- h) Draw energy level diagrams to illustrate the above two reactions.



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