13 Al -> "e + 26 Ma 48K -> = e + 40 G balance nuclear reaction equations aluminum-26 undergoes positron emission potassium-40 undergoes beta decay (electron) 238 () -> 4 He + 234 Th uranium-238 undergoes alpha decay electron configurations 152 25226 35236 451345 Write the electron configuration for C. Write the electron configuration for Cr. Write the shorthand notation for Bi. De] 65244145 110603 NOMENCLATURE [naming compounds, writing formulas] Write the formula for each of the following compounds: a. Sulfur hexafluoride SFS b. Lithium nitride Link c. Chromium (III) carbonate Cr2 ((0))3 d. Tin (II) chloride Sn Ce2 e. Ammonium acetate NHUCZHOR (OF CHOCOONHY) f. Mercury (I) chloride Hazelz g. Potassium bromate KBrog h. Hydrosulfuric acid H2 Some i. Chloric acid Sulfurous acid 42300 Name each of the following compounds. a Cuso, copper (II) sulfate aluminum flooride b. AIF. c. HI d. NO narosen monovale c. H₂Sc (40) hydroselen ic acid f. HNO, notric acid narous acrd g. HNO: NaHSO_a Soulden

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g? 40 -> 20 -> 10 -> 5 -> 2.5 -> 1.25 ->

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

3 half lives 100-750-25-3 [2.5 mg]

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65,

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

63
$$\times .692 = 43.596 \rightarrow 49$$

65 $\times .309 = 20.02 \rightarrow 20.$
 $\frac{+}{63.62} \rightarrow 64$

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D Kirk

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Table of Contents Merrill Chemistry Chapter Review Answers

- 1. Understanding the eBook Merrill Chemistry Chapter Review Answers
 - The Rise of Digital Reading Merrill Chemistry Chapter Review Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Merrill Chemistry Chapter Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Merrill Chemistry Chapter Review Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Merrill Chemistry Chapter Review Answers
 - Personalized Recommendations
 - $\circ\,$ Merrill Chemistry Chapter Review Answers User Reviews and Ratings
 - Merrill Chemistry Chapter Review Answers and Bestseller Lists
- 5. Accessing Merrill Chemistry Chapter Review Answers Free and Paid eBooks
 - Merrill Chemistry Chapter Review Answers Public Domain eBooks
 - Merrill Chemistry Chapter Review Answers eBook Subscription Services
 - Merrill Chemistry Chapter Review Answers Budget-Friendly Options

- 6. Navigating Merrill Chemistry Chapter Review Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Merrill Chemistry Chapter Review Answers Compatibility with Devices
 - Merrill Chemistry Chapter Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Merrill Chemistry Chapter Review Answers
 - Highlighting and Note-Taking Merrill Chemistry Chapter Review Answers
 - Interactive Elements Merrill Chemistry Chapter Review Answers
- 8. Staying Engaged with Merrill Chemistry Chapter Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Merrill Chemistry Chapter Review Answers
- 9. Balancing eBooks and Physical Books Merrill Chemistry Chapter Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Merrill Chemistry Chapter Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Merrill Chemistry Chapter Review Answers
 - Setting Reading Goals Merrill Chemistry Chapter Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Merrill Chemistry Chapter Review Answers
 - Fact-Checking eBook Content of Merrill Chemistry Chapter Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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