PART A —CHEMISTRY

1. An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism:
   (1) tertiary alcohol by S_N1
   (2) secondary alcohol by S_N2
   (3) tertiary alcohol by S_N2
   (4) secondary alcohol by S_N1

Sol. (1)
Reaction proceeds through carbocation formation as 3° carbocation is highly stable, hence reaction proceeds through S_N1 with 3° alcohol.

2. The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of Na^+ will be:
   (1) – 5.1 eV
   (3) –10.2 eV
   (3) + 2.55 eV
   (4) – 2.55 eV

Sol. (1)
$Na \xrightarrow[\Delta H = -5.1 eV]{\Delta H = +5.1 eV} Na^+ + e^-$, here the backward reaction releases same amount of energy and known as Electron gain enthalpy.

3. Stability of the species Li_2, Li_2^- and Li_2^+ increases in the order of:
   (1) $Li_2^- < Li_2^+ < Li_2$
   (3) $Li_2 < Li_2^- < Li_2^+$
   (3) $Li_2^- < Li_2 < Li_2^+$
   (4) $Li_2 < Li_2^+ < Li_2^-$

Sol. (1)
$Li_2 (6) = \sigma 1s^2 \; \acute{\sigma}1s^2 \; \sigma 2s^2$

B.O. $= \dfrac{4-2}{2} = 1$

$Li_2^+ (5) = \sigma 1s^2 \; \acute{\sigma}1s^2 \sigma 2s^1$

B. O. $= \dfrac{3-2}{2} = 0.5$

$Li_2^- (7) = \sigma 1s^2 \; \acute{\sigma}1s^2 \sigma 2s^2 \; \acute{\sigma}2s^1$

B.O. $= \dfrac{4-3}{2} = 0.5$

Li_2^+ is more stable than Li_2^- because Li_2^- has more numbers of antibonding electrons.

4. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
   (1) 1.00 M
   (2) 1. 75 M
   (3) 0.975 M
   (4) 0.875 M

Sol. (4)
$M_1V_1 + M_2V_2 = MV$

$M = \dfrac{M_1V_1 + M_2V_2}{V} = \dfrac{0.5 \times 750 + 2 \times 250}{1000}$

$M = 0.875$

5. Which of the following is the wrong statement?
   (1) O_3 molecule is bent
   (2) Ozone is violet-black in solid state
   (3) Ozone is diamagnetic gas
   (4) ONCl and ONO^- are not isoelectronic

Sol. (All the options are correct statements)
   (1) Correct, as is bent.

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