

# JEE Main 2020 Paper



**Sol. 2**



P - OH bonds = 4

P = O bonds = 2

P - O - P linkage = 1

Ans. 4, 2, 1

**21.** The mole fraction of glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) in an aqueous binary solution is 0.1. The mass percentage of water in it, to the nearest integer, is \_\_\_\_\_.

**Sol. 47 %**

$$x_{\text{Glucose}} = 0.1$$

$$\begin{aligned} \text{mass\% of glucose} &= \frac{0.1 \times 180}{0.1 \times 180 + 0.9 \times 18} \times 100 \\ &= \frac{1800}{18 + 16.2} \\ &= \frac{1800}{34.2} \\ &= 52.63\% \\ &= 53\% \end{aligned}$$

$$\therefore \text{mass \% of H}_2\text{O} = 47\%$$

**22.** The volume strength of 8.9 M  $\text{H}_2\text{O}_2$  solution calculated at 273 K and 1 atm is \_\_\_\_\_. ( $R = 0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}$ ) (rounded off to the nearest integer)

**Sol. 100**

$$\begin{aligned} \text{Vol. strength} &= \frac{8.9}{2} \times \frac{0.0821 \times 273}{1} \\ &= 99.73 \\ &= 100 \end{aligned}$$

**23.** An element with molar mass  $2.7 \times 10^{-2} \text{ kg mol}^{-1}$  forms a cubic unit cell with edge length 405 pm. If its density is  $2.7 \times 10^3 \text{ kg m}^{-3}$ , the radius of the element is approximately \_\_\_\_\_  $\times 10^{-12} \text{ m}$  (to the nearest integer).

**Sol. 143**

$$\text{Density} = \frac{Z \times \text{GMM}}{N_A \times a^3}$$

$$\begin{aligned} 2.7 \times 10^3 &= \frac{Z \times 2.7 \times 10^{-2}}{6.023 \times 10^{23} \times (405 \times 10^{-12})^3} \\ Z &= 6.023 \times 405 \times 405 \times 405 \times 10^{23-36+3+2} \\ Z &= 6.023 \times 405 \times 405 \times 405 \times 10^{-8} \\ Z &= 4 \end{aligned}$$

FCC

$$4R = \sqrt{2} \times a$$

$$R = \frac{405}{2\sqrt{2}} \times 10^{-12} = 143.21 \times 10^{-12} \text{ m}$$

= 143 ans

# Jee Main Paper In 6 4 2014

**JS Bruner**



## **Jee Main Paper In 6 4 2014:**

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