JEE Main 2020 Paper



Sol. 2

- 21. The mole fraction of glucose (C₆H₁₁O₆) in an aqueous binary solution is 0.1. The mass
- percentage of water in it, to the nearest integer, is _____.
 Sol. 47 %

$$x_{\text{Olement}} = 0.1$$
mass% of glucose = $\frac{0.1 \times 180}{0.1 \times 180 + 0.9 \times 18} \times 100$
= $\frac{1800}{18 + 16.2}$
= $\frac{1800}{34.2}\%$
= 52.63%
= 53%

- ... mass % of H₂O = 47%

 22. The volume strength of 8.9 M H₂O₂ solution calculated at 273 K and 1 atm is ______

 (R = 0.0821 L atm K³ mol⁻¹) (rounded off of the nearest integer)
- Sol. 100

Vol. strength =
$$\frac{8.9}{2} \times \frac{0.821 \times 273}{1}$$

= 99.73
= 100

 $Z \times GMM$

- 23. An element with molar mass 2.7x 10° kg mol° forms a cubic unit cell with edge length 405 pm. If its density is 2.7x 10° kg m°, the radius of the element is approximately x 10° m (to the nearest integer).
- Sol. 143

- 143 ans

Density =
$$\frac{Z \times 2.7 \times 10^{-3}}{N_A \times 3^3}$$

2.7 × 10³ = $\frac{Z \times 2.7 \times 10^{-3}}{6.023 \times 10^{23} \times (405 \times 10^{-12})^3}$
Z = 6.023 × 405 × 405 × 405 × 10²³⁻³⁶⁻³⁺³
Z = 6.023 × 405 × 405 × 405 × 10⁻⁶
Z = 4

FCC

4R = $\sqrt{2} \times 3$

R = $\frac{405}{2\sqrt{2}} \times 10^{-12}$ = 143.21 × 10⁻⁶²m

Jee Mains 2question Paper Solution By Fiit Jeee

S. Chand Experts

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