

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

O_2 is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.84 \text{ g H}_2\text{O}$$

c) determine the number of grams of H_2O produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{5 \text{ mol O}_2} \times \frac{44.09 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = .948 \text{ g C}_3\text{H}_8 \text{ used}$$

$$14.8 \text{ g C}_3\text{H}_8 - 0.948 \text{ g C}_3\text{H}_8 = 13.9 \text{ g C}_3\text{H}_8 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{294.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{24.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

Limiting Reactions 12 3 Answers

Morris Hein, Susan Arena, Cary Willard



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Holt Chemistry Ralph Thomas Myers,2004 **Chemical Principles, Properties, and Reactions** Kenneth L. Henold,Frank Walmsley,1984 **Anesthesiology Review: A Comprehensive Q&A Guide** Ivan

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Basic Chemistry Steven S Zumdahl, 1996 *Journal of the American Medical Association* American Medical Association, 1925 *Teaching Science for Understanding* James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities Sunday School Times , 1893

SCIENCE AND ART DEPARTMENT THE COMMITTEE OF COUNCIL ON EDUCATION , 1874
DIRECTORY WITH REGULATIONS FOR ESTABLISHING AND CONDUCTING SCIENCE SCHOOLS & CLASSES , 1871
Directory, with regulations for establishing and conducting science and art schools and classes Education Ministry of, 1900

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Table of Contents **Limiting Reactions 12 3 Answers**

1. Understanding the eBook **Limiting Reactions 12 3 Answers**
 - The Rise of Digital Reading **Limiting Reactions 12 3 Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Limiting Reactions 12 3 Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Limiting Reactions 12 3 Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Limiting Reactions 12 3 Answers**
 - Personalized Recommendations
 - **Limiting Reactions 12 3 Answers** User Reviews and Ratings
 - **Limiting Reactions 12 3 Answers** and Bestseller Lists

5. Accessing Limiting Reactions 12 3 Answers Free and Paid eBooks
 - Limiting Reactions 12 3 Answers Public Domain eBooks
 - Limiting Reactions 12 3 Answers eBook Subscription Services
 - Limiting Reactions 12 3 Answers Budget-Friendly Options
6. Navigating Limiting Reactions 12 3 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactions 12 3 Answers Compatibility with Devices
 - Limiting Reactions 12 3 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactions 12 3 Answers
 - Highlighting and Note-Taking Limiting Reactions 12 3 Answers
 - Interactive Elements Limiting Reactions 12 3 Answers
8. Staying Engaged with Limiting Reactions 12 3 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactions 12 3 Answers
9. Balancing eBooks and Physical Books Limiting Reactions 12 3 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactions 12 3 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactions 12 3 Answers
 - Setting Reading Goals Limiting Reactions 12 3 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactions 12 3 Answers
 - Fact-Checking eBook Content of Limiting Reactions 12 3 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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