

Honors Gifted Chemistry Mole Unit Review Sheet ANSWER KEY

Directions: Read the following questions carefully. SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER. No work, no credit, no kidding. Remember to use significant figures where necessary as well as units and substances (or lose points). If dimensional analysis is appropriate for a problem and it is not used, you will not receive credit for the answer. Place answer in the blank space to the far right where appropriate. Good luck!

1. In lab, a group measures out 102.06 g of peroxide (H_2O_2) and 116.88 g of salt (NaCl). The students in the lab group begin discussing the specific amount in terms of moles and particles. What claims could you make in terms of the number of moles and particles present in each sample? Explain using CER. Show any calculations needed which could provide evidence!

The key to understanding this question is that students are discussing the **AMOUNT** of substances measured in lab. "Amount" refers to **MOLES** since a mole is defined as the SI unit to measure amounts (large amounts of small particles). This means you need to find the number of moles students measured out for each substance by converting the mass GIVEN (in grams) to **MOLES** using the molar mass of each substance.

$$\frac{102.06 \text{ g H}_2\text{O}_2}{1} \times \frac{1 \text{ mole H}_2\text{O}_2}{34.02 \text{ g H}_2\text{O}_2} = 3.0000 \text{ mol H}_2\text{O}_2$$

$$\frac{116.88 \text{ g NaCl}}{1} \times \frac{1 \text{ mole NaCl}}{58.44 \text{ g NaCl}} = 2.0000 \text{ mol NaCl}$$

- This work above means that if students massed out 102.06 g of peroxide, then the sample of peroxide contains 3.0000 moles of peroxide.
- If the student massed out 116.88 g salt, then the sample of salt contains 2.0000 moles of salt.
- Although the mass of the salt is greater than the mass of the peroxide (116.88 g vs 102.06 g), there are more moles of the peroxide present (3.0000 mol) compared to the smaller moles of salt (2.0000 mol). This is due to the smaller molar mass of peroxide.
- Further, if we know the moles of each substance, we can calculate the number of molecules of peroxide and formula units of salt using the fact that 1 mol of anything = 6.02×10^{23} particles.

$$\frac{102.06 \text{ g H}_2\text{O}_2}{1} \times \frac{1 \text{ mole H}_2\text{O}_2}{34.02 \text{ g H}_2\text{O}_2} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}_2}{1 \text{ mole H}_2\text{O}_2} = 1.806 \times 10^{24} \text{ molecules H}_2\text{O}_2$$

$$\frac{116.88 \text{ g NaCl}}{1} \times \frac{1 \text{ mole NaCl}}{58.44 \text{ g NaCl}} \times \frac{6.02 \times 10^{23} \text{ fu NaCl}}{1 \text{ mole NaCl}} = 1.204 \times 10^{24} \text{ fu NaCl}$$

- This work above means that if students massed out 102.06 g of peroxide vs 116.88 g salt, there would be more particles of the peroxide (1.806×10^{24}) than the salt (1.204×10^{24}).
2. The molar mass of hydrogen is 2.02 g/mol. What conversion factors can be written using this information?
 $2.02 \text{ g H}_2 = 1 \text{ mol H}_2$ and this means that $1 \text{ mol H}_2 = 6.02 \times 10^{23} \text{ molecules H}_2$
3. Examine the names of the compounds below.
- Write the chemical formula for each substance and place it in the blank.
 - Use the box to calculate the molar mass of each of the compounds. Show all work!
 - Circle the name of the compound with the **LARGEST** molar mass.

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