

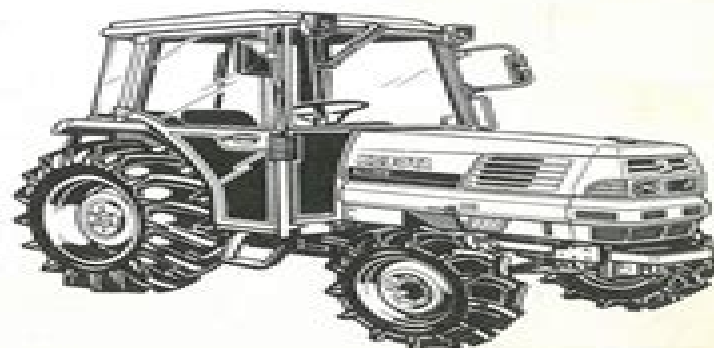
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web in this lab you will collect a measurable volume of butane gas using water displacement using dalton s law and the ideal gas law the number of moles of butane gas and then the molar mass of butane can be calculated your experimental value for molar mass will be compared to the known molar mass of butane and percent error calculated

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web use the combined gas law to determine the volume in l of butane at stp use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 l at stp experimental molar mass g mole of butane the molecular formula of butane is c 4 h 10

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration