

La Bikina

♩ = 160

Words & Music by Ruben Fuentes Gasson
Arranged by Rafael Reis

A Intro

1 $E\flat$ maj7 $A\flat$ maj7 G^7 Cm^7 $D\flat^7(\text{add}13)$ $E\flat$ maj7 Cm^7 Fm^7 $B\flat^7$ $E\flat$ maj7

Se-ñor

B

6 $E\flat$ maj7 $Dm^7(\flat 9)$ $G^7(\flat 13)$ Cm^7 $B\flat m^7$ $E\flat^7$

ta - ri a comi - na la bi - ki - na La gen - te se po - ne a mu - tu a ra -
ne - ra, pre - cio sa ya or - gu - llo - sa No per mite la que ran con - so lar -

10 $A\flat$ maj7 G^7 1. Cm^7 Cm^7 $F^7(\text{add}9)$ Cm^7 $F^7(\text{add}9)$

Di - cen que tie - ne u na pe - na Di - cen que tie - ne u na pe - na que ha ce llo -
Pa - sa lu - cie - do sa re -

14 $Fm^7(\text{add}9)$ $B\flat^7(\flat 13)$ 2. Cm^7 $D\flat^7(\text{add}13)$ $E\flat$ maj7 Cm^7 Fm^7 $B\flat^7$ $E\flat$ maj7

rar Al - ta - al ma je - tad Pa se, ca - mí na, los mí ra sin ver los ja - más

18 $A\flat m^7$ $D\flat^7$ $G\flat$ maj7 $E\flat m^7$ $A\flat m^7$ 1. $D\flat^7$ $G\flat$ maj7 $G^7(\text{add}13)$ 2. $F^7(\text{add}9)$ Fm^7 $B\flat^7$

La bi - ki - na Tie - ne pe - na y - do - lor no co - me - ce eb - mor Al - te

D

24 $E\flat$ maj7 $Dm^7(\flat 9)$ $G^7(\flat 13)$ Cm^7 $B\flat m^7$ $E\flat^7$

ne - ra, pre - cio sa ya or - gu - llo - sa No per mite la que ran con so lar -

28 $A\flat$ maj7 G^7 Cm^7 $D\flat^7(\text{add}13)$ $E\flat$ maj7 Cm^7 Fm^7 $B\flat^7$ $E\flat$ maj7

Di - cen que al - guie - na vi - no y se fue Di - cen que pa - sa las no - ches llo ran do por él

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate BrO_3^- ions

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