

KINETIC THEORY OF GASES

SYNOPSIS

1. BOYLE'S LAW:

- It states that at constant temperature the volume of a given mass of ideal gas is inversely proportional to its pressure.

$$V \propto 1/P \text{ or } PV = \text{constant}$$

$$\text{or } P_1 V_1 = P_2 V_2 \text{ (For constant } m \text{ and } T)$$

2. CHARLE'S LAW:

- It states that if the pressure remains constant, then the volume of a given mass of a ideal gas increases or decreases by $1/273.15$ of its volume at 0°C for each 1°C rise or fall of temperature. Mathematically,

$$V_t = V_0 \left(1 + \frac{t}{273.15} \right) = V_0 \left(\frac{273.15 + t}{273.15} \right)$$

If we put $273.15 + t = T$ and $273.15 = T_0$, then

$$\frac{V_t}{T} = \frac{V_0}{T_0} \text{ or } \frac{V}{T} = \text{constant}$$

So charles' law can also be stated as follows:
At constant pressure, the volume of a given mass of a gas is directly proportional to its absolute temperature.

3. GAYLUSSAC'S LAW:

- It states that if the volume remains constant, then the pressure of the given mass of a gas increases or decreases by $1/273.15$ of its pressure at 0°C for each 1°C rise or fall of temperature. Mathematically,

$$P_t = P_0 \left(1 + \frac{t}{273.15} \right) = P_0 \left(\frac{273.15 + t}{273.15} \right)$$

$$\text{or } \frac{P_t}{P_0} = \frac{T}{T_0} \text{ or } \frac{P}{T} = \text{constant}$$

So Gay Lussac's law can also be stated as

follows: At constant volume, the pressure of a given mass of a ideal gas is directly proportional to its absolute temperature.

4. IDEAL GAS EQUATION:

- For n moles of a ideal gas,

$$PV = nRT \text{ or } \frac{PV_1}{T_1} = \frac{PV_2}{T_2}$$

For 1 mole of a gas

$$PV = RT$$

5. Universal gas constant:

- It signifies the work done by the gas per mole per kelvin. $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1} = 1.98 \text{ cal/mole} \cdot \text{K}$

6. Boltzmann's constant:

- It is the gas constant per molecule of a gas.

If N_A is Avogadro's number, then

$$k_B = \frac{R}{N_A} = 1.38 \times 10^{-23} \text{ J molecule}^{-1} \text{ K}^{-1}.$$

7. IDEAL OR PERFECT GAS:

- A gas which obeys gas laws strictly is an ideal or perfect gas. The molecules of such a gas are of point size and there is no force of attraction between them. The actual real gases obey the ideal gas equation only approximately at low pressures and high temperatures.

8. ASSUMPTIONS OF KINETIC THEORY OF GASES:

- All gases consist of molecules. The molecules are rigid, elastic spheres identical in all respects for a given gas and different for different gases.
- The size of a molecule is negligible compared with the average distance between two molecules.
- The molecules are in a state of continuous random motion, moving in all directions with all possible velocities.
- During the random motion, the molecules collide with one another and with the walls of the vessel.
- The collisions are perfectly elastic and there are no forces of attraction or repulsion between them.
- Between two collisions, a molecule moves in a straight path with a uniform velocity.

Kinetic Theory Reimforecent Section 1 Wprksheet Answer

Christian Drosten



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