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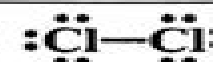
This Exam Is Composed of Three Exercises. It Is Inscribed on Two Pages, Numbered 1 and 2.
Answer the Following Three Exercises:

Exercise 1 (7 points)

Chlorine Gas

Chlorine gas (Cl_2) is very toxic at room temperature. Mixing bleaches containing chlorine element such as Javel water with acidic solutions such as vinegar produces chlorine gas. This gas attacks the respiratory system and may lead to death.

Document-1 shows the Lewis electron dot structure of the chlorine gas molecule.



Document-1

- Pick out from the text the reason why Javel water and vinegar should never be mixed.
- Referring to **Document-1**, answer the following questions:
 - Identify the type of bond in the chlorine gas molecule.
 - Show that chlorine atom has 7 valence electrons.
 - Choose, from the following, the electron configuration of the chlorine atom.
Given: Chlorine element belongs to row 3 (period 3) in the periodic table.
a) K^2L^7 b) $\text{K}^2\text{L}^8\text{M}^7$ c) $\text{K}^2\text{M}^8\text{L}^7$
 - Deduce the atomic number of the chlorine element.
- Chlorine exists in nature as a mixture of different atoms.

Document-2 is a table that shows these chlorine atoms and their percentage abundance in nature.

Referring to **Document-2**, answer the following questions:

- Justify that ^{35}Cl and ^{37}Cl are isotopes.
- Answer by true or false. Correct the false statements.
 - The atoms ^{35}Cl and ^{37}Cl have the same valence.
 - The atom ^{35}Cl has 35 neutrons.
 - ^{35}Cl is less abundant in nature than ^{37}Cl .

Atom	% abundance
^{35}Cl	x%
^{37}Cl	25%

Document-2

Exercise 2 (6 points)

Lithium-Ion Batteries

Lithium-ion batteries are rechargeable cells that use lithium ion as a key component. These batteries deliver a current of high intensity. They are used in portable electronic devices such as smartphones. They are also used to power electrical systems.

Document-1 is a table that shows the composition of a lithium ion.

Number of Protons	Number of Electrons	Number of Neutrons
3	2	4

Document-1

- Pick out from the text one use of lithium-ion batteries.
- During the functioning of such batteries, lithium atoms (Li) at the anode lose their valence electrons.
Referring to **Document-1**, answer the following questions:
 - Verify that the relative charge of the lithium ion is 1+.

Given:

- Relative charge of an ion = relative charge of nucleus + relative charge of electron cloud.
- Relative charge of a proton = 1+; relative charge of an electron = 1-

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Karin Nielsen-Saines



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