

## Section 5.2 Review

### The Structure of the Atom

#### Multiple Choice

For each question below, select the letter of the best answer.

- Which of the following statements about protons, neutrons, and electrons is *true*?
  - Protons and electrons have about the same mass.
  - Neutrons and electrons have equal but opposite charges.
  - The mass of a proton is nearly 2000 times greater than the mass of an electron.
  - The mass of an proton is nearly 2000 times greater than the mass of a neutron.
  - Neutrons and electrons are found in the nucleus of an atom.
- What is the atomic number of mercury, Hg?
  - 2+
  - 80
  - 120
  - 200
  - 200.6
- What is the mass number of an atom of chlorine with 18 neutrons?
  - 1-
  - 17
  - 18
  - 35
  - 35.5
- What element is represented by the Bohr-Rutherford diagram below?
  - potassium, K
  - calcium, Ca
  - hydrogen, H
  - argon, Ar
  - yttrium, Y
- Which of the following statements about the masses of atoms is *false*?
  - Most of the mass of an atom is located in the nucleus.
  - Electrons contribute very little to the mass of an atom.
  - Atoms of different elements have different masses.
  - Most of the mass of an atom's nucleus is due to the mass of neutrons.
  - The mass number of an atom is equal to the number of protons plus the number of neutrons.
- A neutral atom has eight electrons and eight neutrons. Which of the following options correctly represents the atom?
  - carbon-16
  - oxygen-16
  - oxygen-24
  - oxygen-8
  - sulfur-8
- Which of the following symbols represents the element with an atomic number of 14?
  - C
  - N
  - Ni
  - Si
  - S
- Which of the following statements is *false*?
  - The number of neutrons an atom has defines what element it is.
  - The number of electrons an atom has defines what element it is.
  - The number of protons an atom has defines what element it is.
  - All of the above are true.
  - One of the above are true.



# Mcgraw Hill Ryerson Science 9 Answers Key

**Rosina Ehmann**



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