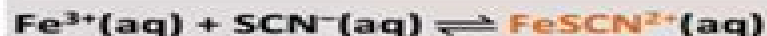


## IA lab work on spectrophotometric determination of $K_c$ of the Iron(3)Thiocyanate system



$$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$$

Since I will not change the temperature in my experiment, the equilibrium constant will stay the same, and therefore I can measure the light absorbance and use the Lambert Beer-Law to find the initial concentration of each species in the system and then calculate the  $K_c$ .

$$X(\text{initial concentration of Iron3 and Thiocyanate}) = A(\text{absorbance of Iron(3)Thiocyanate solution}) / \epsilon \cdot l$$

### Equipment:

- 4 test tubes with 4 corks
- test tube rack
- 10 ml pipette with pump thing

### Solutions:

- 20 ml of Iron (3) solution
- 20 ml of Thiocyanate solution
- 20 ml of water  $\text{H}_2\text{O}$

Samples	The amount of Thiocyanate solution $\text{SCN}^{-}$	The amount of water $\text{H}_2\text{O}$	Iron 3 solution $\text{Fe}^{3+}$
1	2 ml	8 ml	5 ml
2	4 ml	6 ml	5 ml
3	5 ml	5 ml	5 ml
4	9 ml	1 ml	5 ml

# Kc Of Iron Thiocyanate Lab Report

**Sabine Zange**



## Kc Of Iron Thiocyanate Lab Report:

*Chemical Synergies* Nuno A.G. Bandeira, Bartosz Tylkowski, 2018-05-07 This book gives an overview of recent integrated and inter disciplinary approaches between chemical experiment and theory in a variety of fields from polymer science to materials chemistry and ranging from the design of tailored properties to catalysis and reactivity building on the well established success of Density Functional Theory as the foremost quantum chemical method to provide qualitative and quantitative interpretation of results from the chemical laboratory The combination of several characterization techniques with an understanding at the molecular level of chemical and physical phenomena are the main focal point of the subject matter

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