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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	1. 4 molecules	2. 4 molecules NO
1.00 mol	2.00 mol	3. 2.00 mol	4. none	5. none
4.00 mol	6.00 mol	6. 4.00 mol	7. NO	8. 2.00 mol O_2
6.00 mol	1.00 mol	9. 1.00 mol	10. NO	11. 5.00 mol O_2
1.00 mol	6.00 mol	12. 2.00 mol	13. O_2	14. 5.00 mol NO
6.000 mol	6.200 mol	15. 6.200 mol	16. NO	17. 0.400 mol O_2
32.00 g	60.02 g	18. 92.02 g	19. none	20. none
64.00 g	60.02 g	21. 66.01 g	22. O_2	23. 30.12 g NO
64.00 g	30.01 g	24. 28.76 g	25. O_2	26. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 1.302 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of NO/O_2 is larger than the balanced equation mole ratio of NO/O_2 , there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

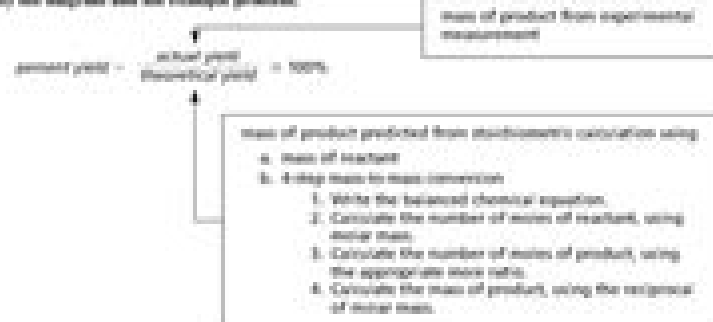
Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g Ga} \times \frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} = 1 \text{ mol Ga}$ 7.00 g Ga \times 1 mol Ga/69.72 g Ga
- number of moles of product: $0.100 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}} = 0.100 \text{ mol } (\text{Ga}_2\text{O}_3)$ 0.100 mol Ga \times 2 mol Ga_2O_3 /4 mol Ga
- theoretical yield: $0.100 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{187.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 18.74 \text{ g } (\text{Ga}_2\text{O}_3)$ correct
- percent yield: $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{18.74 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 39.6\%$ 7.42 g Ga_2O_3 /18.74 g Ga_2O_3 \times 100

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Michael Seilmaier



Mcgraw Hill Companies Night Study Guide:

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Table of Contents Mcgraw Hill Companies Night Study Guide

1. Understanding the eBook Mcgraw Hill Companies Night Study Guide
 - The Rise of Digital Reading Mcgraw Hill Companies Night Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Companies Night Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Companies Night Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Companies Night Study Guide
 - Personalized Recommendations
 - Mcgraw Hill Companies Night Study Guide User Reviews and Ratings
 - Mcgraw Hill Companies Night Study Guide and Bestseller Lists

5. Accessing Mcgraw Hill Companies Night Study Guide Free and Paid eBooks
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 - Mcgraw Hill Companies Night Study Guide Budget-Friendly Options
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 - Mcgraw Hill Companies Night Study Guide Enhanced eBook Features
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 - Interactive Elements Mcgraw Hill Companies Night Study Guide
8. Staying Engaged with Mcgraw Hill Companies Night Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Companies Night Study Guide
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Companies Night Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Companies Night Study Guide
 - Setting Reading Goals Mcgraw Hill Companies Night Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Companies Night Study Guide
 - Fact-Checking eBook Content of Mcgraw Hill Companies Night Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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