

CHAPTER **11** STUDY GUIDE

Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

- _____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
- _____ 2. Stoichiometry is based on the law of conservation of mass.
- _____ 3. In any chemical reaction, the mass of the products is less than the mass of the reactants.
- _____ 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
- _____ 5. The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.



Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01			
9. Water	18.02			

10. What are the reactants? _____
11. What are the products? _____
12. What is the total mass of the reactants? _____
13. What is the total mass of the products? _____
14. How do the total masses of the reactants and products compare? _____

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Lingsheng Yao



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Table of Contents Mcgraw Hill Chemistry Study Guide Answer

1. Understanding the eBook Mcgraw Hill Chemistry Study Guide Answer
 - The Rise of Digital Reading Mcgraw Hill Chemistry Study Guide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Chemistry Study Guide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Chemistry Study Guide Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Chemistry Study Guide Answer
 - Personalized Recommendations
 - Mcgraw Hill Chemistry Study Guide Answer User Reviews and Ratings
 - Mcgraw Hill Chemistry Study Guide Answer and Bestseller Lists

5. Accessing Mcgraw Hill Chemistry Study Guide Answer Free and Paid eBooks
 - Mcgraw Hill Chemistry Study Guide Answer Public Domain eBooks
 - Mcgraw Hill Chemistry Study Guide Answer eBook Subscription Services
 - Mcgraw Hill Chemistry Study Guide Answer Budget-Friendly Options
6. Navigating Mcgraw Hill Chemistry Study Guide Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Chemistry Study Guide Answer Compatibility with Devices
 - Mcgraw Hill Chemistry Study Guide Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Chemistry Study Guide Answer
 - Highlighting and Note-Taking Mcgraw Hill Chemistry Study Guide Answer
 - Interactive Elements Mcgraw Hill Chemistry Study Guide Answer
8. Staying Engaged with Mcgraw Hill Chemistry Study Guide Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Chemistry Study Guide Answer
9. Balancing eBooks and Physical Books Mcgraw Hill Chemistry Study Guide Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Chemistry Study Guide Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Chemistry Study Guide Answer
 - Setting Reading Goals Mcgraw Hill Chemistry Study Guide Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Chemistry Study Guide Answer
 - Fact-Checking eBook Content of Mcgraw Hill Chemistry Study Guide Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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