



John Deere Rx7owners Manual

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web oct 30 2023 two handfuls of blueberries or raspberries one heaped tablespoon of dried fruit such as sultanas currants

or raisins three heaped tablespoons of vegetables such as sweetcorn or peas three heaped tablespoons of any pulse beans
 peas or lentils one cereal bowl of raw leafy greens such as lettuce watercress or spinach

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 water $n_a = 54 \text{ g} / 18 \text{ g mol}^{-1} = 3 \text{ mol}$ total number of moles $n_a + n_b = 0.5 + 3 = 3.5 \text{ mol}$ mole fraction of solute ethyl alcohol $x_b = n_b / n_a + n_b = 0.5 / 3.5 = 0.1429$

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 40 g of H_2O molar mass of H_2O is 18 gmol^{-1} and molar mass of CH_3OH is 32 gmol^{-1} solution moles of CH_3OH $4.5 / 32 = 0.14$
 mole moles of H_2O $40 / 18 = 2.2$ moles therefore according to the equation

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 time the total pressure $p_a = x_a p_t$ mole fraction $x_a = \frac{\text{moles of } a}{\text{total moles}}$ mole fraction of substance
 a is often denoted by x_a and

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the o₂ is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of o₂ are mixed in a flask at room temperature

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