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web indicators are used to determine whether a solution is acidic or alkaline acids react with metals bases and carbonates to  
 produce salts neutralisation is the reaction between an acid and a base

**why is a neutralisation reaction exothermic example socratic** - Sep 22 2022

web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic  
 eg 1 hcl aq + naoh aq → nacl aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup> eg 2 h<sub>2</sub>no<sub>3</sub> aq + koh aq → kno<sub>3</sub> aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$  you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts